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Biologically enhanced cathode design for improved capacity and cycle life for lithium-oxygen batteries

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Abstract

Lithium-oxygen batteries have a great potential to enhance the gravimetric energy density of fully packaged batteries by 2–3 times that of lithium-ion cells. Recent studies have focused on finding stable electrolytes to address poor cycling capability and improve practical limitations of current lithium-oxygen batteries. In this study, the catalyst electrode, where discharge products are deposited and decomposed, was investigated since it plays a critical role in the operation of rechargeable lithium-oxygen batteries. Here we report the electrode design principle to improve specific capacity and cycling performance of lithium-oxygen batteries by utilizing high efficiency nanocatalysts assembled by M13 virus with earth abundant elements, such as manganese oxides. By incorporating only 3–5 wt % of palladium nanoparticles in the electrode, this hybrid nanocatalyst achieves 13,350 mAh g\textsuperscript{−1} (7,340 mAh g\textsuperscript{−1 c+catalyst}) of specific capacity at 0.4 A g\textsuperscript{−1} and a stable cycle life up to 50 cycles (4,000 mAh g\textsuperscript{−1} c+catalyst) at 1 A g\textsuperscript{−1}.

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Author contributions
D.O., J.Q., Y.S.H. and A.M.B. conceived and designed the experiments. D.O., J.Q. performed experiments and analyzed the data. Y.C.L. and Y.S.H. analyzed the electrochemistry data. Y.Z. conducted high-resolution transmission electron microscopy. Y.S.H. and A.M.B. supervised the project. All authors discussed the results and commented on the manuscript.

Additional information
Supplementary information is available in the online version of the paper. Reprints and permission information is available online at http://www.nature.com/reprints.

Competing financial interests
The authors declare no competing financial interests.
Introduction

Rechargeable Li-O₂ batteries operate through the reversible reaction between lithium ions and oxygen molecules. When Li-O₂ batteries discharge in non-aqueous electrolytes, Li⁺ reacts with reduced oxygen (oxygen reduction reaction; ORR) yielding the reaction product, mainly Li₂O₂ that is deposited on the catalyst electrode at a thermodynamic voltage of 2.96 V (vs. Li/Li⁺). Upon charging, the electrical energy is stored by decomposing Li₂O₂ back into Li⁺ and O₂ (oxygen evolution reaction, OER). To promote these ORR and OER for Li-O₂ batteries, several materials have been investigated as catalyst electrodes, such as noble metals (e.g. Au, Pt, Ru, Pd), transition metal oxides (e.g. α-MnO₂, Co₃O₄, MnCo₂O₄, RuO₂), and carbon-based materials (e.g. graphene, carbon fiber, doped carbon nanotube). However, these materials suffer either from high material costs, low conductivities or side reactions with electrolytes, which limits their use as high performance Li-O₂ battery electrodes. Although these difficulties restrain the capability of Li-O₂ batteries, the rational design and selection of electrode materials may fully deliver these battery chemistries to a high gravimetric energy storage system.

To enhance structural functionality, organisms in nature form skeletal tissues from soft to hard through the interactions between nucleating protein matrices and inorganic ions. Inspired by these biominerals researchers have fabricated novel nanoarchitectures using peptides or proteins as scaffolds applied in various functional structures. For example, numerous types of platforms such as synthetic peptide nanofibers, nanorings, and ferritin cages provide simple and uniform ways to synthesize interesting materials at the nanoscale. In electrochemical energy storage devices, nanostructured materials enhance Li ion batteries by shortening the diffusion length of Li ions and benefit capacitors by providing electrodes with large surface areas. For Li-O₂ batteries, since the oxygen is in principle derived from the air, both the oxygen flow into electrodes and the high efficiency to promote discharge/charge reactions are additional important factors. Constructing a new catalyst that combines porous nanostructures and high catalytic activity would allow us to harness the advantages of Li-O₂ batteries for practical applications, such as long-range electric vehicles.

Here we built the catalyst structures to achieve high Li-O₂ battery performances by forming a nanocomposite of biotemplated manganese oxide nanowires (bio MO nanowires) with incorporation of a small weigh percent (3 wt %) of Pd nanoparticles (Fig. 1). We believe this is the first bio-directed synthetic method demonstrated for Li-O₂ battery applications.

Results

M13 virus mediated synthesis of manganese oxide nanowires

Bio MO nanowires with spherulitic surface morphology were synthesized through a facile and moderate chemical reaction by utilizing the biomolecule, M13 virus. M13 viruses have been used as versatile templates for high aspect ratio nanowire synthesis (e.g. semiconductors, metal oxides) under environmentally benign conditions. Among the five different types of capsid proteins on this virus template, p8 is a major coat protein with 2,700 copies surrounding the single-stranded DNA. In addition to the material specificity of these surface proteins (p8 or five copies of minor proteins existing at one end of virus particle, p3), electrostatic interactions between the opposite charge of precursor ions and functional groups of p8 coat protein broaden the template functionality. In this work, the virus clone with the p8 peptide sequences ADVYESALPDPAEAAFE (named FC#2) was used because it has two additional negative functional groups (-COOH from glutamic acid on 13th and 17th) than M13KE (New England Biolabs) under basic conditions. Using this clone, bio MO nanowires were synthesized by first binding Mn²⁺ ions to the FC#2 virus p8
proteins followed by reacting with KMnO₄ at room temperature. From this simple aqueous biotemplating reaction, homogeneous bio MO nanowires with high aspect ratio (~80 nm in diameter and ~1 μm in length) were obtained. These nanowires showed spherulitic surface morphology (TEM image in Fig. 2a, b and Supplementary Fig. S1a). While the flower-like spheres of MO were previously synthesized, the nucleation of MO along the virus particle developed this spherulitic surface of materials into higher dimensional nanowires. Thus, the hierarchical nanostructures of these bio MO nanowires are expected to be advantageous in Li-O₂ battery performances. While previous studies utilized smooth surface MO nanowires for Li-O₂ battery electrodes, the rough surface of these bio MO nanowires can provide large catalytic area as well as enough storage space for discharge products. To the best of our knowledge, this is the first report of M13 virus directed synthesis of manganese oxide nanowires, which is applicable not only to Li-O₂ battery electrodes but also to various applications such as electrochemical capacitors, Li ion battery electrodes and water purifications.

With the aqueous, room temperature approach using M13 virus templates, multi-valent (Mn³⁺/Mn⁴⁺), birnessite-like MO nanowires were developed. The crystallographic and chemical properties of bio MO nanowires were investigated with X-ray diffraction (XRD), high-resolution transmission electron microscopy (HRTEM), the chemical titration method, and X-ray photoelectron spectroscopy (XPS). First, the XRD pattern of virus-templated MO nanowires showed three broad peaks centered at 2θ of 37°, 53.5° and 65.5° (Supplementary Fig. S1b), partially matching to δ-MnO₂ (birnessite, space group C2/m, powder diffraction file (PDF) no. 01-073-2509) XRD pattern. Also, the surface morphology of bio MO nanowires was similar to that of typical birnessites, as can be seen in the TEM image (Fig. 2b). The crystal sizes of bio MO nanowires were 2–5 nm and the co-existing amorphous phase was observed by HRTEM (Fig. 2c). The average oxidation state of bio MO nanowires was determined to be 3.51 by chemical titration methods, suggesting a multi-valence state (Mn³⁺/Mn⁴⁺ = 0.96) of nanowires. This mixed-valence property of MO was also observed from XPS spectra in Supplementary Fig. S2. We confirmed that the average oxidation state of bio MO nanowires lies between Mn⁵⁺ and Mn⁶⁺ by observing two features, the peak position of Mn 2p₃/₂ and the multiplet splitting of Mn 3s. First, the Mn 2p₃/₂ peak position of bio MO nanowires (642.1 eV in Supplementary Fig. S2a) was close to that of MnO₂ (Mn⁴⁺, 642.2 ~ 642.4 eV) as reported in the literature. However, the width of Mn 3s multiplet splitting (5.2 eV in Supplementary Fig. S2b) of bio MO nanowires showed a larger separation than that of MnO₂ (4.5 ~ 4.7 eV) but close to that of Mn₂O₃ (5.2 ~ 5.4 eV). Since it is known that the separation between Mn 3s splitting peaks becomes larger as the oxidation state of Mn is lower, the larger separations of Mn 3s splitting compared to Mn⁴⁺ imply the presence of a lower valence state in bio MO nanowires.

**Nanocompositing ORR catalyst on bio MO nanowires**

To enhance the ORR activity and surface conductivity of porous bio MO nanowires for Li-O₂ battery electrodes, 1–3 wt % of ORR catalytic metal nanoparticles were homogeneously incorporated onto the surface of the nanowires. It is reported that noble metal nanoparticles (e.g. Pd, Pt, Ru and Au) have higher intrinsic ORR activities compared to carbon with 0.1 M lithium perchlorate (LiClO₄) in dimethoxyethane (DME) electrolyte. In particular, very recent studies reported a great stability (Au) and high ORR catalytic activity (Pd) in Li-O₂ battery applications. However, those electrodes contained 40–100 wt % of noble metal nanoparticles, making them less practical due to cost efficiency. Thus, we suggest the composite structure of bio MO nanowires combined with noble metal nanoparticles to decrease the loading of precious metals in cathodes while maximizing the catalytic activity of Li-O₂ batteries.
The nanocomposites were designed to nucleate the metal nanoparticle on the surface of the bio MO nanowire templates to maximize the contact with discharge products in Fig. 1 (red lines, right middle). In addition, the homogeneous distribution of ORR metal catalyst along the bio MO nanowires was important in maintaining the porous structure of the electrodes. To facilitate these requirements, poly acrylic acid (PAA) was incorporated around the bio MO nanowires to nucleate the metal nanoparticle. The PAA wrapped nanowires increased the colloidal stability of the nanowires during the synthesis so that a more homogeneous solution of bio MO nanowires was observed (Supplementary Fig. S3). Control experiment without PAA wrapping resulted in free Pd nanoparticle formation and aggregated MO nanowires as shown in the TEM image (Supplementary Fig. S4). This composite synthesis method was applied to two different ORR metal catalysts, Au or Pd (Au/Pd). To nucleate these metal nanoparticles on PAA wrapped bio MO nanowires, ethylene glycol was used as a mild and selective reducing agent to react with noble metal precursors (HAuCl₄, Na₂PdCl₄) slowly, avoiding reactions with virus-templated MO nanowires. This metal nanoparticle incorporation method by stabilizing and trapping precursor ions with PAA is expected to be applicable not only to biotemplated nanowires but also to various nanosized substrates, such as nanofibers.

The homogeneous distribution of Au/Pd nanoparticles (with a diameter of 2–6 nm, Supplementary Fig. S5) was observed by elemental mapping using scanning transmission electron microscopy (STEM) in Fig. 2d, e. The amount of PAA used to form the composite nanostructure was determined with thermogravimetric analysis (TGA) to be 2.5 wt % of the electrode by comparing the weight loss difference between bio MO nanowires and Pd incorporated bio MO nanowires at 500°C (Supplementary Fig. S6). The amount of Au/Pd nanoparticle incorporated in the bio MO nanowires was measured by inductively coupled plasma atomic emission spectroscopy (ICP-AES) analysis and varied approximately 1–3 wt % of the total electrode, depending on the metal precursor amount that was added during synthesis.

**Improved Li-O₂ battery performances with bio MO nanowires**

The virus-templated MO Li-O₂ battery cathodes showed a 9,196 mAh g⁻¹ (955 mAh g⁻¹+catalyst+Li₂O₂) first discharge capacity with 92.6 % of coulombic efficiency at a current density of 0.4 A g⁻¹. The electrochemical tests were performed with 0.1 M LiClO₄/DME electrolyte under 1 atm of oxygen (further details on measurement methods are described in the experimental section). Although ether-based electrolytes may undergo some side reactions during Li-O₂ battery charging, it is one of the most stable electrolytes against the oxygen electrode for Li-O₂ batteries known to date without the reactivity with Li metal like DMSO has. Further control experiments with different electrolytes were included in the methods section and Supplementary Fig. S7. For galvanostatic tests, the operating voltage was limited to 2.2–4.15 V (vs. Li/Li⁺) to minimize possible parasitic reactions. First, the effect of MO electrode structures on Li-O₂ battery performance was studied by varying the architecture from nanoparticles to bio MO nanowires. Under the same synthetic conditions, MO nanoparticles (60 nm diameter) were synthesized without viruses and their morphology was compared to bio MO nanowires with scanning electron microscope (SEM) images (Fig. 2f, g). The surface area of both the nanoparticles and nanowires was measured by the Brunauer-Emmett-Teller (BET) method giving 98.5 m² g⁻¹ and 271.7 m² g⁻¹, respectively. Bio MO nanowires showed a 40 % improvement of Li₂O₂ storage capability (9,196 mAh g⁻¹ at 0.4 A g⁻¹) compared to MO nanoparticles (6,545 mAh g⁻¹), as can be seen from the first discharge capacity in Fig. 3a. In addition, the discharge voltage of the bio MO nanowires (~2.68 V) is higher than that of MO nanoparticles (~2.6 V) by ~80 mV and the charge voltage of the bio MO nanowires (~3.65 V) is lower than that of MO nanoparticles (~3.77 V) by ~120 mV. In other words, the bio MO nanowires showed
decreased discharge and charge overpotentials compared to the MO nanoparticles. Since the oxidation state of both the MO nanoparticles and nanowires was determined to be similar by XPS analysis (Supplementary Fig. S2a, b), the difference in Li-O_2 battery discharge capacity was mainly attributed to their different surface area and morphologies. Since MO nanoparticles agglomerate during the synthesis, the low surface area electrode materials (Fig. 2f) limit the surface mediated reaction so that it results a poor specific discharge capacity. In our biotemplated system, the high aspect ratio as well as larger surface areas (Fig. 2g) formed highly active Li-O_2 battery cathodes where the electrode spaces can be fully utilized while minimizing the effect of discharge product blocking the electrodes.

The major discharge product obtained after the first discharge of bio MO nanowire electrodes was Li_2O_2 confirmed by XRD in Fig. 3b. The XRD data of discharged electrode matches well to the reference XRD pattern of Li_2O_2 (orange star, PDF no. 01-074-0115). There was no other major discharge product observed except Li_2O_2. In previous reports, MOs have been shown to produce a LiOH phase during the discharge due to the structural water or hydroxyl residues. The Li-O_2 battery performance of dehydrated (thermal treatments at 160°C for 2 hours) bio MO nanowires was investigated to eliminate any possible contributions from water molecules or hydroxyl groups. The thermally treated nanowires showed a similar discharge capacity of 9,206 mAh g\(^{-1}\) at 0.4 A g\(^{-1}\) (Supplementary Fig. S7, blue) compared to nanowires without heat treatments, although structural water decreased from 3.2 to 1.5 wt % of electrodes after heat treatment (Supplementary Fig. S6). Thus, the LiOH formation is negligible in this Li-O_2 battery performance.

**Battery performances of Pd/Au loaded bio MO nanowires**

The hybrid nanostructure composed of Pd nanoparticles and bio MO nanowires exhibited improvement of the Li-O_2 battery discharge capacity. The first galvanostatic cycle of Au/bio MO nanowires and Pd/bio MO nanowires were compared with approximately the same amount of Au (1.7 wt %) and Pd (1.1 wt %) incorporations in each nanocomposited electrode. After the addition of only ~1 wt % of Au or Pd nanoparticles, the first discharge capacities of Au/bio MO nanowires and Pd/bio MO nanowires (cut off voltage at 2.2 V vs Li/Li\(^+\)) were improved by 8 % (Au) and 15.5 % (Pd) respectively, compared to bare bio MO nanowires (at 0.4 A g\(^{-1}\)). Au: blue straight, Pd: red dashed). In addition, the Pd/bio MO nanowire electrodes showed approximately 60 mV higher discharge voltage than Au/bio MO nanowires as can be observed from Fig. 3c. The higher discharge voltage of Pd/bio MO nanowires than Au/bio MO nanowires in the first galvanostatic cycle of Li-O_2 electrodes may be attributed to the attractive interaction between Pd and oxygen. Although Pd has a lower contrast than Au on TEM observation due to its lower atomic mass, large (higher than 5 nm) Pd nanoparticles on bio MO nanowires were rarely observed than Au nanoparticles decorated bio MO nanowires (Supplementary Fig. S5) so the smaller nanoparticle distribution of Pd/bio MO nanowires in a given synthetic conditions may improve the battery performances. Thus, the surface electrocatalytic functionality of bio MO nanowires was easily tuned by the composite synthesis method developed here.

The capability to store lithium peroxide was further improved by increasing Pd incorporations from 1.1 wt % to 3 wt % of Pd/bio MO nanowires electrodes. With 3 wt % of Pd incorporation, a 45 % improvement of the first specific capacity of bio MO nanowires (Fig. 3c, red straight, 13,347 mAh g\(^{-1}\), 1,008 mAh g\(^{-1}\) c+ catalyst + Li_2O_2) was achieved at 0.4 A g\(^{-1}\). This is the highest specific capacity under a high applied current (0.4 A g\(^{-1}\), 0.22 A g\(^{-1}\) c+ catalyst) among all manganese-oxide-based electrodes reported to date. The formation of Li_2O_2 after discharge and its decomposition during the charge (oxygen evolution was detected by Gas Chromatography (GC) in Supplementary Table. S1) with Pd/bio MO nanowire electrodes was also confirmed by Raman spectroscopy in Supplementary
Fig. S8. No significant amount of side product (e.g., Li$_2$CO$_3$) can be detected by XRD and Raman, suggesting that DME is largely stable during discharge.$^{54}$ We believe the enhanced discharge capacity of Pd/bio MO nanowires originated from the better utilization of porous electrode spaces by improving the electrochemical active surface area and surface conductivity of bio MO nanowires. The Pd nanoparticles located on the surface of the bio MO nanowires facilitated ORR reactions and improved the electronic conduction throughout the Li$_2$O$_2$ layers allowing further discharge product depositions.

The differences in the charging profile of the 3 wt % Pd-MO nanowires compared to the bare bio MO nanowires (Fig. 3a) can be associated with the increased thickness of Li$_2$O$_2$ layer of the 3 wt % Pd/bio MO nanowires (~4 nm Li$_2$O$_2$) compared to the bare bio MO nanowires (~2.6 nm Li$_2$O$_2$) due to larger discharge capacity. However, the charging potential of Pd/bio MO nanowires after decomposing 0.05 μAh cm$^{-2}$true of Li$_2$O$_2$ was still ~100 mV lower than that of catalyst-free, carbon electrode$^{55}$ at the same depth of discharge, ~1 μAh cm$^{-2}$ (the true electrode surface area specific capacities, the detail true calculation method is included in the methods section). This lower OER potential compared with catalyst-free electrodes suggests the catalytic function of Pd/bio MO nanowires even after the deep discharge in Li-O$_2$ battery. The following (second) cycle profile of MO nanoparticles, bio MO nanowires, Au (1.7 wt%)/bio MO nanowires, Pd (1.1 wt%)/bio MO nanowires and Pd (3 wt%)/bio MO nanowires are included in supplementary Fig. S9 and Table. S2. The second charging profile showed enhanced performance than previous manganese-oxides-based electrode at ~6 times higher current density$^{38}$ and further studies to improve the cycle life after the deep discharge remains as future work.

To clarify the role of Pd nanoparticles, PAA wrapped bio MO nanowires were tested as Li-O$_2$ battery electrodes under the same test conditions. The first discharge capacity of the PAA/bio MO nanowires electrode was 9,722 mAh g$^{-1}$c (Supplementary Fig. S7, green), which is only a 5 % improvement of bio MO nanowire electrodes. This small improvement of discharge capacity could be attributed to the decreased aggregation between bio MO nanowires after PAA stabilization. The average value of discharge specific capacity can be seen in Fig. 3d. Therefore, the increased first discharge capacity of Pd/bio MO nanowire composites originated from Pd nanoparticle incorporation onto bio MO nanowires, not from the PAA wrapping. We expect that this methodology can be easily applied to other promising ORR or OER candidate materials to develop efficient catalyst structures for Li-O$_2$ batteries in future studies.

**Lowered carbon mass in cycling electrodes**

The catalyst nanostructure designed in this work, Pd/bio MO nanowires, allowed for an electrode with greatly decreased carbon content (8 wt %) in the catalyst electrodes to improve cycle life of Li-O$_2$ battery. Although carbon functions as a conducting agent and an ORR catalyst, it can also reduce the cycle life because side products can be formed at the carbon/electrolyte or carbon/Li$_2$O$_2$ interface in electrodes with ether-based electrolytes.$^{56}$ By reducing carbon amounts in our electrodes, cycle life of Li-O$_2$ batteries was enhanced at high current densities (1 A g$^{-1}$c; 0.1 A g$^{-1}$c+catalyst) even under relatively low charging cutoff voltage (4.15 V). The cycling tests were conducted by limiting the discharge capacity to 2,000 mAh g$^{-1}$c (200 mAh g$^{-1}$c+catalyst) or 4,000 mAh g$^{-1}$c (400 mAh g$^{-1}$c+catalyst) following the previous works method.$^{3, 10}$ Similar to the first discharge specific capacity, the cycle life of Li-O$_2$ batteries was gradually enhanced as the catalyst structure changed from MO nanoparticles to nanowires and finally to composite Pd/bio MO nanowires (Fig. 3e, 3f, Supplementary Figs. S10–12). In particular, the OER overpotential was lowered as the surface area of catalyst electrodes as well as the surface conductivity of bio MO nanowires was improved by Pd incorporation (Fig. 3e) after being discharge to 4,000 mAh g$^{-1}$c (400 mAh g$^{-1}$c+catalyst). Thus, the first cycling round trip
efficiencies of MO nanoparticles, bio MO nanowires and Pd/bio MO nanowires was improved from 75%, 77.7% to 79.2%, respectively, as the cathode was rationally designed. The aggregated microstructure of MO nanoparticles limited the Li-O2 battery cycle life to 10 cycles with 4,000 mAh g\(^{-1}\) (299 mAh g\(^{-1}\)catalyst+LiO\(_2\)) at 1 A g\(^{-1}\)c (0.1 A g\(^{-1}\)catalyst). The bio MO nanowires showed cycle retentions up to 20 cycles due to the high porosity and large catalytic surface area of the electrodes. After incorporating Pd nanoparticles, the discharge capacity of Li-O2 batteries maintained stable to 50 cycles (Fig. 3f, Supplementary Fig. S10). When the depth of discharge capacity was fixed to 2,000 mAh g\(^{-1}\)c (at 1 A g\(^{-1}\)catalyst, 0.1 A g\(^{-1}\)catalyst), these electrodes showed a longer cycle life with a similar cycling performance trend. Each MO nanoparticles, bio MO nanowires and Pd/bio MO maintained the 2,000 mAh g\(^{-1}\)c of discharge capacity up to 22, 47 and 58 cycles, respectively (Fig. 3f, Supplementary Fig. S11). The reversible cycling number with Pd/bio MO nanowires electrodes is smaller than the recently reported, which might be due to the instability of electrolyte or the reactivity between carbon and oxygen reduction intermediates. This can be observed from the charging cycle profile in Supplementary Figs. S10–12, resulting in the faster decay of charging capacity than the discharging capacity as cycling number increased in Fig. 3f. Ongoing and future work will be focused on developing stable electrolytes and carbon-free electrodes to enhance the cycling of these electrodes.

To investigate the effect of carbon in cycle of Li-O2 batteries, electrodes were prepared with 44 wt % of carbon. The cycle life of bio MO nanowires and Pd/bio MO nanowires decayed faster than those of low carbon electrodes under the same fixed discharge capacity (728 mAh g\(^{-1}\)c, 400 mAh g\(^{-1}\)catalyst) and current density (1 A g\(^{-1}\)catalyst, 0.1 A g\(^{-1}\)catalyst). The observed cycle number was 17, 21 and 30 for MO nanoparticles, bio MO nanowires and Pd/bio MO nanowires, respectively (Supplementary Fig. S12). Although the electronic conductivity was higher for 44 wt % versus 8 wt % carbon electrodes, side products can be formed in the high carbon electrodes during cycling counteracting the kinetic advantage in Li-O2 battery performance. Therefore, we believe Pd/bio MO nanowires are a viable materials platform for long lasting, high capacity Li-O2 batteries.

**Discussions**

Using a biological template, a new catalyst electrode was developed that increased capacity (13,350 mAh g\(^{-1}\)catalyst, 7,340 mAh g\(^{-1}\)catalyst at 0.4 A g\(^{-1}\)catalyst) and cycle life (50 cycles with 4,000 mAh g\(^{-1}\)catalyst, 400 mAh g\(^{-1}\)catalyst at 1 A g\(^{-1}\)catalyst, 0.4 A g\(^{-1}\)catalyst) of Li-O2 batteries at the highest gravimetric current density compared with other transition metal oxide based electrodes in literatures. These templates allowed the formation of high aspect ratio manganese oxide nanowires wrapped in 3 wt % Pd nanoparticles facilitating compositional control that outperforms materials made by mechanical mixing. These high aspect ratio bio MO nanowires formed a porous network that maximizes the interaction between the catalyst and the discharge products (Li2O2) thus increasing reversibility and generating a stable cycle life of Li-O2 batteries. Furthermore, the Pd/bio MO nanowires allowed the implementation of low carbon electrodes batteries improving battery cycle life. We believe this template is readily adaptable for other materials combinations and can provide a platform to test various catalytic electrode materials for Li-O2 batteries and other catalytic systems. In addition, the application of biomaterials will broaden the material selection scope into diverse areas in designing Li-O2 battery electrodes thus contributing to the development of energy storage systems.
**Methods**

**Materials synthesis**

Manganese sulfate monohydrate (MnSO₄·H₂O, Sigma-Aldrich, 2.4 ml of 50 mM) and sodium sulfate (Na₂SO₄, Sigma-Aldrich, 2.4 ml of 50 mM) solutions were added together into 50 ml of D.I. water. FC#2 viruses (1.63×10¹³) were added into this solution and stirred at room temperature overnight. After the overnight incubation, potassium permanganate (KMnO₄, Alfa Aesar, 16 ml of 10 mM) solution was added and stirred at room temperature overnight. The final products were washed with D.I. water by centrifugation and collected after lyophilization. To incorporate the noble metal nanoparticles (Pd, Au), the bio MO solution was dialyzed against D.I. water overnight and diluted with D.I. water to make a final solution volume of 200 ml. PAA (MW: 2,000, 0.6 g in 20 ml of D.I. water) was added and stirred overnight. The PAA wrapped bio MO nanowires were collected by centrifugation and dispersed again with 240 ml of D.I. water. Au/Pd precursor solutions were added into the previous solution (For Au: gold (III) chloride trihydrate, HAuCl₄·3H₂O, 2.17 ml of 25.4 mM for 1.7 wt % of Au in the final Au/bio MO nanowire composite electrodes. For Pd: sodium tetrachloropalladate (II), Na₂PdCl₄, Sigma-Aldrich, 0.612 ml of 34 mM for 3 wt % (0.408 ml of 34 mM for 1.1 wt %) of Pd in the final Pd/bio MO nanowire composite electrodes) and stirred overnight at room temperature. To reduce the metal ion, 3 ml of ethylene glycol was added to each Au/bio MO nanowire and Pd/bio MO nanowire solution for 24 hours inside a brown bottle due to light sensitivity. The solution was washed three times by centrifugation with D.I. water and lyophilized.

**Materials characterizations**

For SEM images, FEG-SEM 6700F (JEOL) was used under 5 kV with 5.9 mm of working distance (or 0.5 kV with 3 mm of WD) after coating the sample with Pd. High-resolution transmission electron microscopy (HRTEM) images were obtained with 2100, 2010 FEG (JEOL) under an accelerating voltage of 200 kV. The sample was loaded on formvar/carbon or quantifoil grids (Electron microscopy sciences). X-Ray Diffraction (PANanalytical Multipurpose Diffractometer, Cu Kα radiation) patterns were collected with 0.18 ° min⁻¹ scan speeds under the following settings: 2° of anti-scatter slit, 6 mm of irradiated length of automatic mode and 0.04 rad of soller slit. The discharged/charged samples were collected after disassembling batteries inside an Ar-filled glove box and were covered by Kapton® tape to prevent any air contamination. Inductively coupled plasma atomic emission spectroscopy (ICP-AES, Horiba Jobin Yvon ACTIVA-S) was used to determine the atomic ratio between manganese, gold, palladium ions of Au/bio MO nanowires and Pd/bio MO nanowires. The sample was dissolved in nitric acid at 50 °C overnight and diluted with D.I. water (ACS reagent grades) to make the final nitric acid concentration of 2 % v/v for ICP measurements. X-ray photoelectron spectroscopy (XPS, Physical Electronics Versaprobe II) was measured with the monochromatic Al X-ray source (Kα, 1486.6 eV) under the pass energy of 23.5 eV and the step size of 0.1 eV. The acquired XPS spectra were calibrated with adventitious carbon peak (1s) as a reference positioned at 284.8 eV. To quantify the PAA mass of the metal nanoparticles/bio MO nanowires composites, thermogravimetric analysis (TGA, Q50 TA instrument) was used with 10 °C min⁻¹ increasing temperature rate under nitrogen.

**Chemical titration for Mn oxidation state determination**

The average oxidation state of MO was determined by two separate chemical titration steps reported in the literature.⁴³ The color change indicator (sodium diphenylamine-4-sulfonate, C₁₂H₁₀NNaO₃S, Alfa Aesar) was further added during the titration. MO nanoparticles were used for titrations because the manganese oxidation state is expected to be similar to bio MO.
nanowires from XPS analysis and the protein residue of bio MO nanowires would cause significant errors during the titration.

**Electrochemical tests**

The lyophilized catalyst materials were further dried in a vacuum oven at 90 °C overnight before electrode preparations. The Li-O₂ electrodes were composed of Ketjan black (EC600JD, AkzoNobel), catalyst materials, LiTHion™ dispersion (Ion Power, USA) with the mass ratio of 44: 36: 20 for high carbon containing (44 wt %) electrodes and the mass ratio of 8: 72: 20 for low carbon containing (8 wt %) electrodes. These three components were dissolved into 2-propanol and mixed overnight using a bath sonicator to get homogeneous slurry mixtures. The mixture was spread on the separator (Celgard, PP2075) by tape casting methods and dried in a vacuum oven at 80°C for 2 hours. The dried slurry was punched into circle electrodes with diameters of 1.27 cm and the mass of each electrode was measured (0.1–0.3 mg cm⁻²) following further drying in vacuum oven at 80°C overnight before the electrochemical test. Inside the Ar-filled glove box, the catalyst electrode was soaked with the electrolyte in advance (0.1 M LiClO₄ in DME, Novolyte) and assembled into the Li-O₂ battery cell designed by our lab. Metallic Li foils (negative electrode), two layers of separators, catalyst electrodes and a positive electrode current collector (316 stainless steel mesh) were sequentially put on the cell and additional electrolyte (100 μl) was added. After the assembly, the Li-O₂ battery cell was purged with oxygen (Airgas, UPC grade) and closed at 1.08 atm. The purged cell was allowed to rest at least 2 hours before the electrochemical test. The first cycle of Li-O₂ battery cell was tested by galvanostatic measurements at 0.4 A g⁻¹c with a Solartron Analytical 1470E potentiostat. For the cycling tests, the pressure of the Li-O₂ battery cell was adjusted again to 1.08 atm after 16 cycles and 1 A g⁻¹c of current was applied.

**The Li-O₂ battery performance with different electrolytes**

The first galvanostatic cycle of bio MO nanowires was tested with 1 M lithium triflate (LiCF₃SO₃) in tetra (ethylene) glycol dimethyl ether (TEGDME) to investigate the Li-O₂ performance dependence on electrolytes. Here, bio MO nanowire electrodes with LiCF₃SO₃/TEGDME exhibited higher charging overpotential as well as lower first discharge capacity (Supplementary Fig. S7, red. 8,766 mAh g⁻¹c at 0.4 A g⁻¹c). Furthermore, the coulombic efficiency of nanowire electrodes dropped from 92.6 % (0.1 M LiClO₄ in DME) to 61.5 % (1 M LiCF₃SO₃ in TEGDME) at the same charging cut off voltage, 4.15 V. Since increasing the test voltage windows higher than 4.15 V may evoke the electrolyte decomposition, DME was suitable for bio MO nanowire electrodes in order to achieve higher coulombic efficiency in lower voltage ranges.

**The true electrode surface area specific capacity**

Here we make estimation in calculating the thickness of the discharge product, Li₂O₂. It is assumed that Li₂O₂ is formed as films throughout the overall catalyst electrode with the volumetric capacity, 2698.6 mAh cm⁻³. First, the true electrode surface area was calculated. The BET surface area of bio MO nanowires is 271.7 m² g⁻¹ and the reported BET surface area of Ketjan black (EC600JD, AkzoNobel) is 1,400 m² g⁻¹. Thus, the true surface area can be obtained by considering the weight percentage (catalyst: Ketjan black = 36: 44) of each component in electrode, resulting in 271.7 m² g⁻¹ * 0.36 +1,400 m² g⁻¹ * 0.44 = 713.8 mtrue² g⁻¹c+catalyst. As 13,347 mAh g⁻¹c corresponds to 7,340.8 mAh g⁻¹c+catalyst, the true electrode surface area specific capacity is ≈1.0 μAh cm⁻²true.
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References


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Figure 1. Schematic of a nanocomposite structure
Synthesis step of the metal nanoparticle/M13 virus-templated manganese oxide nanowires (bio MO nanowires) and the operational reaction inside Li-O$_2$ battery cells.
Figure 2. Electron microscope images of catalyst materials

TEM images of M13 virus-templated MO nanowires (a) at low magnitude (scale bar corresponds to 500 nm), (b) zoomed at the end of an individual bio MO nanowire (scale bar, 50 nm) and (c) HRTEM image showing amorphous phase mixed with crystalline phase (scale bar, 5 nm, Inset: the Fast Fourier Transform (FFT) pattern with two major rings matching to the atomic planes with the spacing of 2.4 Å and 1.6 Å, respectively). The synthetic procedure of bio MO nanowires was easily reproducible. We synthesized, characterized, and made electrodes from over 30 times independent experiments. (d) STEM image of Pd/bio MO nanowires (left) and the corresponding Energy Dispersive X-ray (EDX) elemental mapping images of manganese (Mn, red, upper right) and palladium (Pd, blue, lower right, scale bars, 100 nm). (e) STEM image of Au/bio MO nanowires (left) and the corresponding EDX elemental mapping images of manganese (Mn, red, upper right) and gold (Au, yellow, lower right, scale bars, 50 nm). The SEM image and BET surface area of (f) MO nanoparticle showing bulky phase of aggregated particles and (g) homogeneously dispersed bio MO nanowires with larger pores (scale bars, 500 nm). The two BET surface areas are 98.5 (f) and 271.7 m²g⁻¹ (g).
Figure 3. The electrochemical performance of Li-O$_2$ batteries

(a) The first galvanostatic cycle of MO nanoparticles (blue) and bio MO nanowires (red) at a current density of 0.4 A g$^{-1}$ ($\approx$0.044 mA cm$^{-2}$) with 0.1 M LiClO$_4$/DME electrolytes under 1 atm oxygen. (b) XRD patterns of bio MO nanowire electrodes at the pristine state (blue), the end of the first full discharge until 2.2 V (dark yellow) and the end of the first full charge until 4.15 V (red) under 0.4 A g$^{-1}$ of current density. (c) The first galvanostatic cycle of bio MO nanowires with the similar weight percent of Pd (red dash) and Au (blue straight) nanoparticle composites at the current density of 0.4 A g$^{-1}$. Further improvement was observed with 3 wt % addition of Pd nanoparticles (red straight) onto bio MO nanowires. (d) The specific capacities of MO nanoparticles (orange), bio MO nanowires (blue), bio MO nanowires with PAA wrapping (dark yellow), Pd/bio MO nanowires (red) electrodes including standard deviations error bars (154–854 mAh g$^{-1}$). Each sample type was made and tested at least triplicate. (e) The first cycling voltage profile of Li-O$_2$ batteries with the fixed discharge capacity of 4,000 mAh g$^{-1}$ (400 mAh g$^{-1}$ for MO nanoparticles (blue), bio MO nanowires (dark yellow) and Pd/bio MO nanowires (red) at the current density of 1 A g$^{-1}$. Each electrode was composed of 8 wt % of carbon and was cycled in 2–4.15 V voltage windows. (f) The cycling performance of Li-O$_2$ battery electrodes composed of MO nanoparticles (blue), bio MO nanowires (dark yellow) and Pd/bio MO nanowires (red) at the current density of 1 A g$^{-1}$ (0.1 A g$^{-1}$ with the two different fixed discharge capacities of 2,000 and 4,000 mAh g$^{-1}$ (200, 400 mAh g$^{-1}$) with the two different fixed discharge capacities of 2,000 and 4,000 mAh g$^{-1}$ (200, 400 mAh g$^{-1}$). The circular shape indicates discharge capacities (for 2,000 and 4,000 mAh g$^{-1}$), the triangle indicates charge capacities (until 4.15 V) for 2,000 mAh g$^{-1}$ and the inverted triangle indicates charge capacities for 4,000 mAh g$^{-1}$. Each electrode was composed of 8 wt % carbon. All of electrochemical tests reported in this work were repeated for at least six different devices.