Theoretical study of the ammonia nitridation rate on an Fe (100) surface: A combined density functional theory and kinetic Monte Carlo study


http://dx.doi.org/10.1063/1.4896610
American Institute of Physics (AIP)
Final published version
Tue Oct 16 05:53:51 EDT 2018
http://hdl.handle.net/1721.1/95885
Article is made available in accordance with the publisher’s policy and may be subject to US copyright law. Please refer to the publisher’s site for terms of use.
Theoretical study of the ammonia nitridation rate on an Fe (100) surface: A combined density functional theory and kinetic Monte Carlo study

Sang Chul Yeo,1 Yu Chieh Lo,2 Ju Li,2 and Hyuck Mo Lee1, a)
1Department of Materials Science and Engineering, KAIST, 291 Daehak-ro, Yuseong-gu, Daejeon 305-701, South Korea
2Department of Nuclear Science and Engineering and Department of Materials Science and Engineering, Massachusetts Institute of Technology, Cambridge, Massachusetts 02139, USA

(Received 16 July 2014; accepted 16 September 2014; published online 3 October 2014)

Ammonia (NH3) nitridation on an Fe surface was studied by combining density functional theory (DFT) and kinetic Monte Carlo (kMC) calculations. A DFT calculation was performed to obtain the energy barriers (Ee) of the relevant elementary processes. The full mechanism of the exact reaction path was divided into five steps (adsorption, dissociation, surface migration, penetration, and diffusion) on an Fe (100) surface pre-covered with nitrogen. The energy barrier (Ee) depended on the N surface coverage. The DFT results were subsequently employed as a database for the kMC simulations. We then evaluated the NH3 nitridation rate on the N pre-covered Fe surface. To determine the conditions necessary for a rapid NH3 nitridation rate, the eight reaction events were considered in the kMC simulations: adsorption, desorption, dissociation, reverse dissociation, surface migration, penetration, reverse penetration, and diffusion. This study provides a real-time-scale simulation of NH3 nitridation influenced by nitrogen surface coverage that allowed us to theoretically determine a nitrogen coverage (0.56 ML) suitable for rapid NH3 nitridation. In this way, we were able to reveal the coverage dependence of the nitridation reaction using the combined DFT and kMC simulations. © 2014 AIP Publishing LLC. [http://dx.doi.org/10.1063/1.4896610]

I. INTRODUCTION

Ammonia (NH3) gas nitridation of metals is a commonly used heat treatment for improving surface hardness, mechanical properties, and wear and corrosion resistance as this process incorporates chemically stable nitride compounds to create multi-functional materials for use in various industrial applications.1–7 However, NH3 gas nitridation is a time-consuming process. Various research groups have investigated NH3 nitridation techniques. Xuan et al.8 employed both simulations and experiments to develop a self-consistent diffusion model for simulating atomic nitrogen transport in titanium alloys through the coupling effect of both the initial activated nitrogen concentration and a high-temperature gradient. Other groups have reported stable austenitic structures in medical stainless steels that meet the requirements for surgical implantation.9 To understand the exact NH3 nitridation process, a fundamental theoretical study is needed.10 Specialy, the mechanisms for rapid NH3 nitridation rate have not yet been elucidated. In addition, investigations of catalytic properties on a metal surface must include an examination of the reaction-poisoning or reaction-promoting influence of surface coverage on various reaction mechanisms. The Ertl group reported adsorbate–absorbate interactions based on a statistical analysis of scanning tunneling microscopy results at 300 K.11 Zheng et al. measured the coverage-dependent diffusion coefficient for the surface diffusion of hydrogen on a flat Pt (111) surface over a wide coverage range, from 0.1 to 0.8 ML (ML stands for mono-layer), by linear optical diffraction (LOD).12 In addition, Nobuhara et al. studied the coverage dependence of the H-induced relaxation of a Pd lattice and the corresponding variation in the energy barrier for H adsorption.13 The kinetic Monte Carlo (kMC) simulations offer a suitable method for determining rate conditions and have recently been applied by a number of groups to examine the relative rates of competing reactions on surfaces.14–20

We studied NH3 gas nitridation in real time via kMC simulations. We propose that pre-coverage with nitrogen on an Fe surface may affect the nitridation rate. The reliability of the kMC simulations depends on accurate transition rates, which denote events.21–25 The effect of surface coverage on an NH3 nitridation process can be divided into 5 steps: adsorption, desorption, dissociation, penetration of the dissociated nitrogen, and diffusion into the bulk Fe. To simulate all nitridation reaction steps, we quantitatively obtained the adsorption energy, dissociation barrier, penetration energy barrier for the dissociated nitrogen, and surface migration energy barrier of the nitrogen for various surface coverages (θ = 0.22 ML, 0.56 ML, 0.78 ML, 1.00 ML when θ is coverage).26–28 This paper is organized as follows. Section II presents the computational methods and systems. Section III provides the density functional theory (DFT) calculation and kMC simulation results. Finally, we discuss our findings in terms of how to improve the nitridation rate using the surface coverage effect and state our conclusions in Sec. IV.

II. COMPUTATIONAL DETAILS

We performed calculations using the Vienna ab initio simulation package, a plane-wave code that employs...
projector augmented wave method.\textsuperscript{29–32} The spin-polarized DFT calculation was performed with the generalized gradient approximation for the exchange-correlation functional using the PBE method.\textsuperscript{33} DFT calculations have previously been shown to aid the mapping of the energetics of various reaction paths. To determine the reaction events, we used a $(3 \times 3)$ super-cell containing a five-layer slab with two fixed layers and a 15-Å vacuum. Only the top 3 layers were allowed to relax. We used an energy cutoff of 400 eV and a $(5 \times 5 \times 1)$ Monkhorst-Pack grid for k-point sampling. We obtained a lattice parameter of 2.83 å, a bulk modulus of 1.76 Mbar, a cohesive energy of 5.16 eV, and a magnetic moment of 2.21 $\mu_B$, which constituted the bulk properties of the pure bcc Fe used for ferromagnetic iron. All values were in good agreement with the literature.\textsuperscript{34} The adsorption energy ($E_{ad}$) of the adsorbate, $\text{NH}_x (x = 0–3)$, for each reaction was calculated according to Eq. (1):

$$E_{ad} = E_{x/surface} - E_{surface} - E_{NH_x},$$

in which $E_{x/surface}$, $E_{NH_x}$, and $E_{surface}$ are the total energies of the adsorbed system, the adsorbate, and the clean surface, respectively. A negative adsorption energy corresponds to an energetically favorable adsorption site on the surface. Larger negative values correspond to stronger adsorption. The energy barrier ($E_b$) for each reaction was computed according to Eq. (2):

$$E_b = E_{TS} - E_{IS},$$

in which $E_b$, $E_{TS}$, and $E_{IS}$ are the energy barrier of the elementary step, the total energy of the transition state, and the total energy of the initial state, respectively.

The energy ($\Delta E$) for each reaction was computed according to Eq. (3):

$$\Delta E = E_{FS} - E_{IS},$$

in which $E_{TS}$, $E_{FS}$, and $E_{IS}$ are the total energy of transition state, final state, and initial state, respectively. A negative reaction energy indicates that the process is exothermic, whereas a positive value indicates that the process is endothermic. We employed the climbing image nudged elastic band (CI-NEB) method to construct the minimum energy pathway (MEP).\textsuperscript{35} To examine the effect of surface coverage, we adopted a pre-covered nitrogen atom within a covered Fe (100) $p(3 \times 3)$ surface ($\theta = 0.22$ ML, 0.56 ML, 0.78 ML, 1.00 ML), as depicted in Fig. 1.

FIG. 1. Structural configuration of a $3 \times 3$ super-cell containing five Fe (100) layers. The “h” in the black circle denotes a hollow site, “b” denotes a bridge site, and “t” denotes a top site. The various N pre-covered surfaces were first optimized ($\theta = 0$ ML, 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML). The purple and blue spheres represent the Fe and N atoms and a vacuum of 15 Å.
We employed the kMC algorithm implemented in our C code to investigate the NH$_3$ nitridation rate on a N pre-covered Fe surface. In a kMC simulation, the temporal evolution of the system is described by a stochastic trajectory. The states along this trajectory are connected by processes associated with the following probabilities:

$$r_i = v_i e^{-E_i/k_B T}$$

where $v_i = 2k_B T/h$ is on the order of the atomic vibrational frequency and $E_i$ is the energy barrier associated with event $i$. $T$ is the temperature, $k_B$ is the Boltzmann's constant, and $h$ is the Planck's constant. This form for the rates implies that, in a kMC model, events with a high probability will occur more frequently. At each time step, all rates are calculated for the given system configuration. In general, the atomic vibrational modes at these points are also used to calculate the pre-factor in the kinetic Monte Carlo. For all events in this study, we used a standard value of $10^{13}$ s$^{-1}$ for the atomic pre-factor $\Omega_1$.

To select which event will occur at a given time step, a total transition rate $R$ is constructed as follows:

$$R = \sum_{i=1}^{N} r_i$$

for $N$ possible events in a given system configuration. A uniformly distributed number $\rho_i \in [0, 1)$ is calculated, and an event $j$ is selected to occur at the given time step. All possible events can be grouped together to reduce the search time for the event. The temporal evolution of the system is based on the assumption that the probability of an event

<table>
<thead>
<tr>
<th>$\theta$</th>
<th>IS</th>
<th>TS</th>
<th>FS</th>
</tr>
</thead>
<tbody>
<tr>
<td>0 ML</td>
<td><img src="image1" alt="Image" /></td>
<td><img src="image2" alt="Image" /></td>
<td><img src="image3" alt="Image" /></td>
</tr>
<tr>
<td>0.22 ML</td>
<td><img src="image4" alt="Image" /></td>
<td><img src="image5" alt="Image" /></td>
<td><img src="image6" alt="Image" /></td>
</tr>
<tr>
<td>0.56 ML</td>
<td><img src="image7" alt="Image" /></td>
<td><img src="image8" alt="Image" /></td>
<td><img src="image9" alt="Image" /></td>
</tr>
<tr>
<td>0.78 ML</td>
<td><img src="image10" alt="Image" /></td>
<td><img src="image11" alt="Image" /></td>
<td><img src="image12" alt="Image" /></td>
</tr>
<tr>
<td>1.00 ML</td>
<td><img src="image13" alt="Image" /></td>
<td><img src="image14" alt="Image" /></td>
<td><img src="image15" alt="Image" /></td>
</tr>
</tbody>
</table>
FIG. 3. Energy diagram of both the energy barrier (E_b) and reaction energy (E) for the penetration of dissociated nitrogen at surface coverages of θ = 0 ML, 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML. The black line indicates the energy barrier; the red line indicates the reaction energy.

III. RESULTS AND DISCUSSION

A. DFT calculation of elementary reaction events

In the DFT calculation, we obtained the adsorption energy (E_ad) and energy barrier (E_b) for the relevant elementary processes within the complete NH₃ nitridation mechanism. The reaction was divided into 5 steps (adsorption, dissociation, penetration, surface migration, and diffusion) on an N pre-covered Fe (100) surface. In the first step, adsorption, we investigated the adsorption properties of N and NH₃ on the Fe (100) surface. The results indicated that N adsorbs onto 4-fold hollow sites (E_ad = −1.66 eV). It was previously reported that NH₃ prefers a top site on the Fe atom for adsorption.¹¹ The NH₃ adsorbs onto a top site (E_ad = −0.92 eV) on a clean Fe surface, and the adsorption energy for NH₃ decreases as the surface coverage increases. In this study, energies of −0.72 eV, −0.48 eV, −0.41 eV, and −0.36 eV were obtained at θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML, respectively.

Next, we investigated the effect of surface coverage on the dissociation reaction, which is considered the high-energy barrier for full ammonia (NH₃) dehydrogenation. To consider the nitrogen surface coverage effect at each reaction event throughout the process, we employed DFT calculations on a low index (100) of the Fe surface, which is the most stable surface. In addition, the nitrogen-covered Fe (100) surface was modeled using different surface nitrogen concentrations.

Fig. 2 indicates that the reaction pathway for NH₃ dissociation is divided into three elementary steps. In these steps, the corresponding sequence of intermediate steps for NH₃ dissociation can be written as follows:

\[ \text{NH}_3 + \ast \rightarrow \text{NH}_2^\ast + \text{H}^\ast, \]  
(4)

\[ \text{NH}_2^\ast + \ast \rightarrow \text{NH}^\ast + \text{H}^\ast, \]  
(5)

\[ \text{NH}^\ast + \ast \rightarrow \text{N}^\ast + \text{H}^\ast. \]  
(6)

The * symbols denote an empty site and an adsorbed NH₃ (a = 0–3) intermediate species on the Fe surface. These dehydrogenation steps are highly reducible. According to Fig. 2, the NH₃ dissociation reactions have a rate-determining step (5). Using the DFT calculations, we obtained the energy barrier (E_b) for the NH₃ dissociation. As the surface coverage increased, the dissociation energy barrier increased. Other research groups have reported that an increase in the surface coverage affects the energy barrier for the dissociation reactions of various species.⁴²–⁴⁴ Our results indicated energy barriers of 1.24 eV, 1.43 eV, 2.05 eV, and 2.30 eV at θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML, respectively. In our previous study, the energy barrier results for the reaction step (4) were...
FIG. 5. Reaction pathway of molecular N in N pre-covered bcc-Fe. The black line indicates its behavior at an N coverage of 0.22 ML; the red line indicates its behavior at an N coverage of 0.56 ML; the blue line indicates its behavior at an N coverage of 0.78 ML; and the green line indicates its behavior at an N coverage of 1.00 ML on Fe (100).

used to estimate the catalytic activity of ammonia decomposition on an Fe catalyst. During the surface migration reaction step, the surface migration energy barrier was considered. Because the dissociated nitrogen atom is expected to migrate from one hollow site to the nearest adjacent hollow site, this surface migration step, which also contributed to the surface coverage effect, should be considered. The results indicated that the surface migration energy barrier was dominant over the surface coverage, with values of 0.9 eV, 0.66 eV, 0.33 eV, and 0.33 eV at $\theta = 0.22$ ML, 0.56 ML, 0.78 ML, and 1.00 ML, respectively.

During the penetration reaction step, a nitrogen atom moves from a hollow site on the surface to an octahedral site in the sub-layer. We performed DFT calculations to examine the effect of surface coverage on the penetration of the surface-dissociated nitrogen atom into the sub-surface. The results are presented in Fig. 3, in which the energy barriers ($E_b$) and reaction energies ($\Delta E$) are plotted for the penetration of a dissociated nitrogen at various surface coverages ($\theta = 0$ ML, 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML). Jiang and Carter indicated that a carbon atom diffuses into Fe (100) and Fe (110) with energy barriers of 1.18 and 1.47 eV, respectively. We previously reported that Fe (100) and Fe (110) have the highest energy barrier for nitrogen penetration, in good agreement with the experimental results presented here. As the surface nitrogen coverage increased on the Fe surface, the penetration energy barrier decreased. We obtained the lowest penetration energy barrier for the dissociated nitrogen ($E_b = 0.73$ eV) with a surface coverage of 0.78 ML. However, at full coverage (1.00 ML), the energy barrier for the penetration of the dissociated nitrogen atom increased ($E_b = 1.82$ eV). Based on the structural analysis, we determined that the low-energy barrier for the penetration of the dissociated nitrogen is governed by the surface stability. As the surface coverage increased, the surface symmetry was broken and distortion increased. For example, the lattice constant is 2.83 Å at the bare surface but, at the coverage of 0.78 ML, it changed to 2.80 Å in one direction while it increased up to 2.92 Å in another direction as seen in Fig. 4. Thus, the surface stability was reduced and nitrogen penetration was easily induced. When the surface is fully covered by nitrogen, the lattice constant changes back to the original value and thus the surface symmetry is restored (see Fig. 4 at the full coverage). The penetration of the dissociated nitrogen is impeded again, which we call the blocking effect. We primarily focused on the surface coverage, which affects the NH$_3$ nitridation reactions on the Fe surface. For the diffu-

FIG. 6. The initial configuration of the bcc Fe (100) (10 x 10) with 40 layers and a vacuum level of 4 Å used in the kMC simulation. In this system, the gray, blue, and white spheres represent the Fe, N, and H atom. Non-surface Fe atoms are rendered invisible.
sion reaction step, we constructed an energy diagram for the diffusion of nitrogen into bulk Fe, a process that is generally well characterized. Nitrogen is typically located in the octahedral interstitial sites. Given the energy barrier for nitrogen confirmed in pure bcc-Fe, our result ($E_b = 0.73$ eV) is in good agreement with the experimental results.$^{48, 49}$

We primarily focused on the effect of surface coverage on NH$_3$ nitridation on the Fe surface. The reaction was divided into adsorption, dissociation, surface migration, penetration, and diffusion. For a quantitative analysis of the surface coverage effect, we reinvestigated NH$_3$ nitridation as a function of surface coverage. A schematic diagram for the full NH$_3$ nitridation mechanism is provided in Fig. 5. We described five steps along the reaction pathway.

### B. kMC simulations of ammonia nitridation

We performed kMC simulations to confirm the surface coverage effects on the complete process: adsorption, desorption, dissociation, reverse-dissociation, surface migration, penetration, reverse-penetration, and diffusion reactions as determined from the DFT calculations. For the kMC simulation, we utilized a system model for the initial configuration as depicted in Fig. 6. Fig. 7 provides a possible process for the different reactions that occur during NH$_3$ nitridation, includ-

**TABLE I.** A list of the nitration reaction events, consisting of the adsorption energy, activation energy, and pre-factor for each reaction step. The pre-factor denotes the default value.

<table>
<thead>
<tr>
<th>Event</th>
<th>Elementary reaction</th>
<th>Energy barrier ($E_b$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Adsorption</td>
<td>NH$_3$ + * → NH$_3$*</td>
<td>0 eV</td>
</tr>
<tr>
<td>Desorption</td>
<td>NH$_3$* → NH$_3$ + *</td>
<td>0.22 ML: 0.72 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 0.41 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 2.05 eV</td>
</tr>
<tr>
<td>Dissociation</td>
<td>NH$_3$* → N* + 3H*</td>
<td>0.22 ML: 1.24 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 2.05 eV</td>
</tr>
<tr>
<td>Reverse dissociation</td>
<td>N* + 3H* → NH$_3$*</td>
<td>0.22 ML: 1.28 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 0.84 eV</td>
</tr>
<tr>
<td>Surface migration</td>
<td>N(hollow) → N(hollow)</td>
<td>0.22 ML: 0.9 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 0.33 eV</td>
</tr>
<tr>
<td>Penetration</td>
<td>N(hollow) → N(o-site)</td>
<td>0.22 ML: 1.83 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 0.73 eV</td>
</tr>
<tr>
<td>Reverse penetration</td>
<td>N(o-site) → N(hollow)</td>
<td>0.22 ML: 0.5 eV</td>
</tr>
<tr>
<td></td>
<td></td>
<td>0.78 ML: 0.4 eV</td>
</tr>
<tr>
<td>Diffusion</td>
<td>N(o-site) → N(o-site)</td>
<td>0.73 eV</td>
</tr>
</tbody>
</table>
FIG. 8. The NH₃ nitridation results from the kMC simulation. These figures indicate the degree of nitrogen diffusion in bulk bcc Fe at 5 × 10⁶ step of kMC simulation. Initial conditions were decided by NH₃ contents (20, 30, 40, 50, 60, 70, 80, and 90 NH₃) on the vacuum level. Non-surface Fe atoms are rendered invisible for better visualization.

To obtain an accurate kMC simulation, the reverse reactions must be considered. We considered two reaction events, the reverse reactions of both the dissociation and penetration steps. These reverse dissociation events were determined at 1.28 eV, 1.12 eV, 0.84 eV, and 0.25 eV at θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML, respectively. Finally, the diffusion events were set at 0.73 eV, the energy barrier for nitrogen diffusion in bulk Fe. Because the diffusion events have no effect on N surface coverage on Fe surface.

To obtain an accurate kMC simulation, the reverse reactions must be considered. We considered two reaction events, the reverse reactions of both the dissociation and penetration steps. These reverse dissociation events were determined at 1.28 eV, 1.12 eV, 0.84 eV, and 0.25 eV at θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML. These values were obtained using the previous NH₃ dissociation results and are also the energy barriers for the rate-limiting step in the complete dehydrogenation reaction pathway for NH₃ dissociation, including the reverse reactions. The reverse penetration events were determined by 0.5 eV, 0.45 eV, 0.4 eV, and 1.6 eV at θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML, respectively. Table I lists the energy barriers for the 8 possible reaction events.

These results describe the behavior of the NH₃ nitridation reaction using a kMC simulation. We conclude that the nitrogen surface coverage affects the nitridation rate. In order to verify the degree of nitrogen diffusion in the bulk bcc Fe at 5 × 10⁶ step of kMC simulation. In Fig. 8, the initial conditions were decided by the NH₃ contents. We provided simulation snapshots of the nitridation process involving ammonia based on the nitrogen surface coverage (θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML). As shown in Fig. 9, we could
easily determine that the number of diffused nitrogen atoms is maximized under the initial condition of 50 NH₃ out of 90 sites on the Fe surface. Considering that the real-time atomistic kMC simulations are based on a database of energy barriers (E₆) obtained from the DFT results, we believe that an increase in the NH₃ nitridation rate can be attained by a nitrogen surface coverage of about 0.56 ML on an N pre-covered Fe surface. It is most likely because distortion of surface is optimum at this value of coverage and the blocking effect becomes dominant at a higher coverage.

IV. CONCLUSION

We studied NH₃ gas nitridation using a combination of DFT and kMC simulations to examine the effect of nitrogen surface coverage on an Fe surface. This research provides a new theoretical approach to estimate the nitridation rate. The mechanism used for the DFT calculation was divided into the adsorption, dissociation, surface migration, and penetration properties (i.e., configuration, site preference, and relative stability) and the diffusion. To investigate the effect of nitrogen surface coverage on each reaction step, we modeled a nitrogen-covered Fe (100) surface with various surface nitrogen concentrations (θ = 0.22 ML, 0.56 ML, 0.78 ML, and 1.00 ML). We then calculated the elementary reaction steps following the mechanism. The reactivity of both the adsorption and dissociation steps decreased as the nitrogen surface coverage increased on the Fe (100) surface. We also examined the surface migration and penetration energy barriers. The surface stability was correlated with the surface symmetry along a surface containing an N pre-covered Fe surface. The nitrogen diffusion was calculated in bulk bcc Fe and found to be in good agreement with the experiment.

Using a database of the DFT calculations, we performed a real-time atomistic simulation using kMC simulations supported by a database of energy barriers (E₆) from the DFT results. The reliability of the kMC simulations depends on accurate transition rates, for events. We concluded that the nitridation rate could be improved by the suitable nitrogen coverage (0.56 ML) on the Fe (100) surface.

ACKNOWLEDGMENTS

This research was supported by a grant from the National Research Foundation of Korea (NRF) funded by the Korean government (MEST) (2011-0028612).