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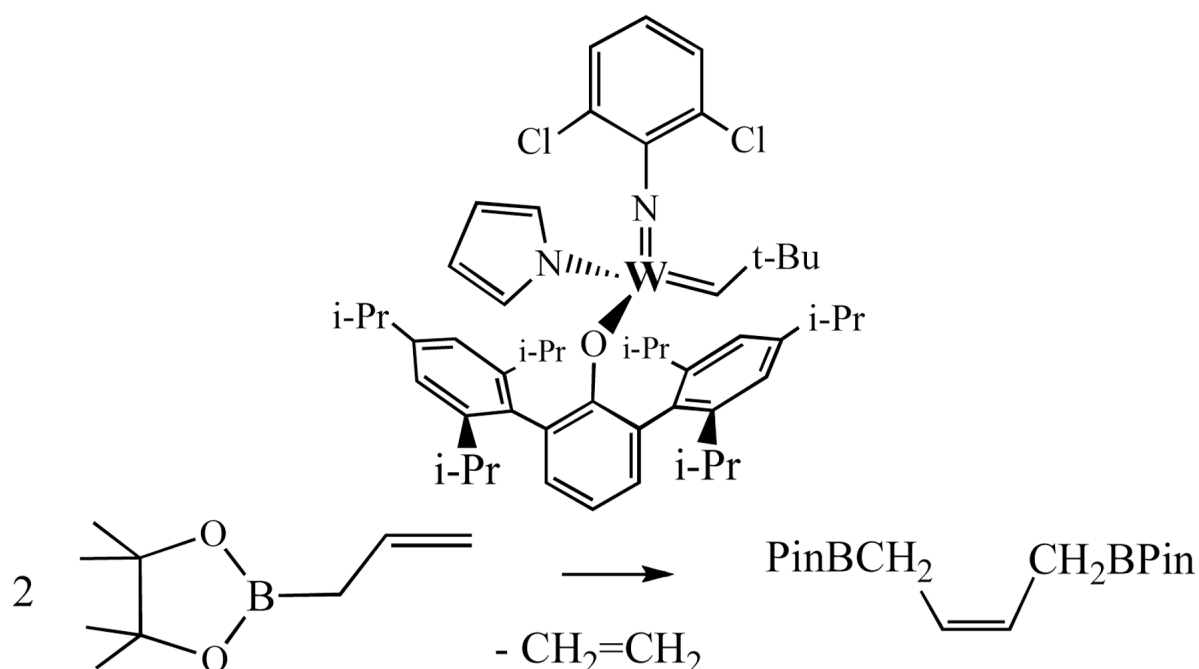
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Highly Z-Selective Metathesis Homocoupling of Terminal Olefins

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Abstract



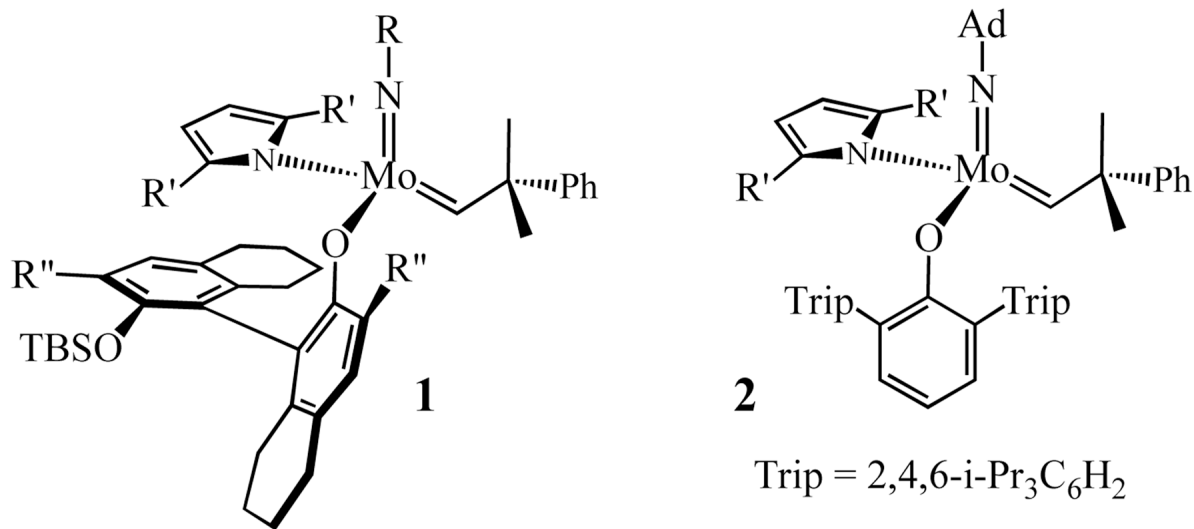
0.2% cat, benzene reflux, 18 h, 74% conversion, 94% Z

Mo and W MonoAryloxy-Pyrrolide (MAP) olefin metathesis catalysts can couple terminal olefins to give as high as >98% Z-products in moderate to high yields with as little as 0.2% catalyst. Results are reported for 1-hexene, 1-octene, allylbenzene, allyltrimethylsilane, methyl-10-undecenoate, methyl-9-decenoate, allylB(pinacolate), allylOBenzyl, allylNHTosyl, and allylNHPh. It is proposed that high Z-selectivity is achieved because a large aryloxy only allows metallacyclobutanes to form that contain adjacent *cis* substituents and because isomerization of Z-product to E-product can be slow in that same steric environment.

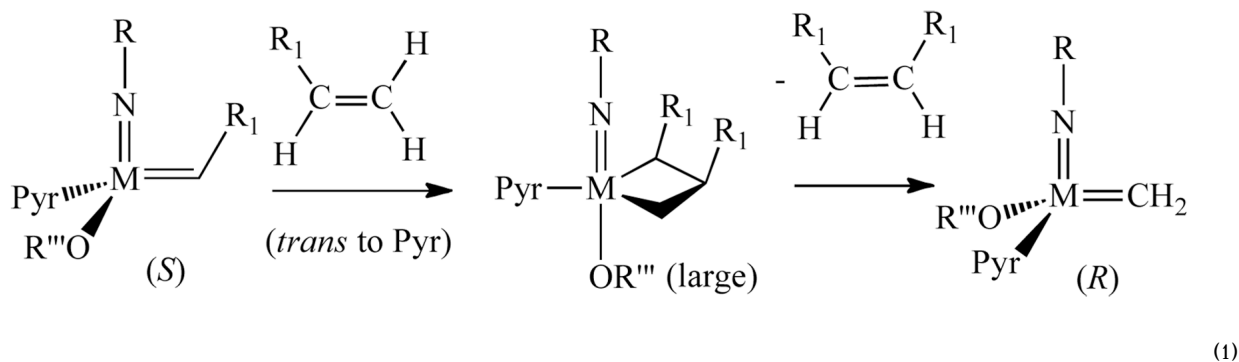
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Supporting Information Available: Experimental details for the synthesis of all compounds and metathesis reactions. Supporting Information is available free of charge via the Internet at <http://pubs.acs.org>.

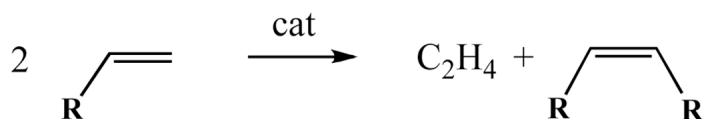
Highly reactive monoaryloxide-pyrrolide (MAP) olefin metathesis catalysts of Mo^I (**1**) have been prepared from bispyrrolide precursors² *in situ* and employed for enantioselective olefin metathesis reactions, such as a total synthesis of quebrachamine (when R = 2,6-*i*-Pr₂C₆H₃, R' = Me, R'' = Br).³ Cross-metatheses with **1** (when R = 1-adamantyl) were found to be *Z*-selective, as well as enantioselective, as a consequence, we proposed, of restrictions imposed on substituents within the metallacyclobutane in the TBP intermediate by a "large" axial aryloxide in combination with a "small" imido group.⁴ The above principle guided the synthesis of related complexes **2** (R' = H (**2_H**) or Me (**2_{Me}**)), which have been shown to catalyze the *Z*-selective metathesis of selected internal *Z*-olefins and to polymerize substituted norbornadienes to yield >99% *cis* and >99% syndiotactic polymers.⁵ We report here that the principles of *Z*-selectivity by MAP catalysts can be extended to homocoupling of terminal olefins.



The principles of *Z*-selective homocoupling of R₁CH=CH₂ by a *syn,rac*-MAP complex are shown in equation 1. (Only *syn* isomers¹ have been observed to date in NMR or X-ray studies.) The terminal olefin enters the coordination sphere *trans* to the pyrrolide (Pyr)^{5,6} to generate an intermediate metallacyclobutane with adjacent R₁ substituents pointing away from the axial OR'' group; this scenario holds *if* OR'' is sufficiently large to prevent formation of any metallacycle in which R₁ is oriented toward OR''. Loss of *Z*-R₁CH=CHR₁ yields a methylene complex with an inverted configuration at the metal center (*S*→*R* in eq 1). A productive metathesis reaction between the methylene species and R₁CH=CH₂ then yields ethylene and reforms (*S*)-M(NR)(CHR₁)(Pyr)(OR''). Inversion of configuration at the metal center is a consequence of each forward metathesis step.^{5,6} Inversion itself should not be an important feature of a homocoupling reaction, but if OR'' is enantiomerically pure, then both diastereomers must be *Z*-selective.



Initially, we concentrated on small-scale experiments involving 1-hexene (S_1) or 1-octene (S_2 ; Table 1; eq 2) with 4 mol % catalyst in a closed system (NMR tube or vial). On the basis of extensive screening (see Supporting Information) we conclude that (i) a relatively small imido group is not necessary for highly *Z*-selective couplings of a *terminal* olefin, and (ii) *W*-based complexes deliver higher %*Z* than *Mo* complexes. Examples of higher selectivity furnished by *W* include 3_W versus 3_{Mo} and 5_W versus 5_{Mo} .⁷ Compound 4_W ^{6a} is as effective as 3_W , and 95% *Z*-product is obtained at low conversion. The *Z*-product isomerizes to *E* with time and conversion (e.g., see 5_W). The results for the homocoupling of 1-hexene are similar to those shown in Table 1 in all cases; for example, catalyst 4_W gave 95% *Z*-5-decene at 33% conversion after three days.



$\text{R} = \text{n-butyl } (S_1), \text{ n-hexyl } (S_2), \text{ CH}_2\text{Ph } (S_3), \text{ CH}_2\text{SiMe}_3 (S_4), (\text{CH}_2)_8\text{CO}_2\text{Me } (S_5),$

$(\text{CH}_2)_7\text{CO}_2\text{Me } (S_6), \text{ CH}_2\text{-B} \begin{array}{c} \diagup \text{O} \diagdown \\ | \quad | \\ \text{C} \quad \text{C} \\ \diagdown \text{O} \diagup \end{array} (S_7), \text{ CH}_2\text{OBenzyl } (S_8), \text{ CH}_2\text{NHTosyl } (S_9), \text{ CH}_2\text{NHPh } (S_{10})$

(2)

A selection of some of our findings in connection with initial screening of substrates S_3 – S_9 is shown in Table 2.⁷ Note that 7_{Mo} and 3_W both give high %*Z* for S_7 , although a direct comparison of *Mo* and *W* is not possible since 1-adamantyl imido species of *W* are not known. It should be noted also that moderate conversion is satisfactory since the product can be separated easily from the unreacted substrate, which can be recycled.

In order to test the degree to which ethylene might be deleterious to *Z*-selectivity, we explored reactions involving several of the higher boiling substrates under a 0.5 or 10 mmHg vacuum with 1 or 2 mol % catalyst on larger scales, and compared the findings with those obtained at 1 atm of nitrogen (Table 3). The results suggest that the effects of carrying out reactions at reduced pressure are not significant. Only the coupling of S_5 carried out with 8_W over 15 h suggests that reduced pressure *may* be required to maintain high %*Z* for long periods. Some loss of monomer at ~0.5 mm naturally is observed over extended times (~14% measured in the case of 8_W after 15 h).

The results of reactions performed at 80–120 °C (bath temp), larger scales, and lower catalyst loadings are shown in Table 4. In several cases, the remaining monomer was removed *in vacuo* and the *Z*-product yields were established. A number of reactions proceed with >90 % *Z*-selectivity and in good yield.

We propose that the efficiency of the mechanism of formation of *Z*-product shown in eq 1 depends upon a ligand combination that allows only a *syn,syn- α ,R₁/ β ,R₁* metallacyclobutane to form from a *syn* alkylidene. Therefore, a large OR''' ligand is required in order to form *Z*-product with high selectivity, as we proposed previously. A "small" imido group is not required, most likely because the steric demands of a *syn,syn- α ,R₁/ β ,R₁* metallacyclobutane are less severe than the steric demands of an all *syn, trisubstituted* metallacycle.⁵

A critical question relates to the mechanism of formation of *E*-product. Three possible direct ways of forming *E*-product are (i) approach of monomer to the *syn* alkylidene to yield a metallacycle with R₁ pointed toward OR'''; (ii) reaction of monomer with a highly reactive (unobservable) *anti* alkylidene (in equilibrium with a *syn* alkylidene) to give a *trans* disubstituted metallacyclobutane intermediate; (iii) approach of the monomer in a manner different from that shown in eq 1 to generate a different type of metallacyclobutane. On the basis of what we know at present we propose that a significant amount of *E*-product is not formed through a direct method if OR''' is sufficiently large (equation 1); formation of *Z*-product in the primary step obviously is required for a *Z*-selective reaction.

One possible indirect mode of forming *E*-product is for the *Z*-product to be isomerized through reaction with a M=CHR₁ species to afford a trisubstituted metallacycle that contains two adjacent *trans* substituents. Such a process is likely to be relatively slow in many circumstances for steric reasons⁵ because two R₁ groups must be oriented toward the large OR''' in the metallacyclobutane, if that metallacyclobutane is the only type that forms. A second possible indirect mode is for the reverse of eq 1 to be fast (ethenolysis⁸), but only if the monomer is reformed and recoupled many times in the presence of ethylene, and if a "mistake" that results in formation of *E*-product in any single step (eq 1 and immediately above) is thereby magnified. On the basis of the results summarized in Table 3, rapid and repeated ethenolysis is probably not the main pathway giving rise to *E*-products with the catalysts and substrates explored here. Therefore, at this stage we propose that *E*-product forms primarily through isomerization of the initial *Z*-product.

Formation of high %*Z* homocoupled acyclic products, as described herein, has, to the best of our knowledge, never been observed. The experiments detailed here have evolved as a consequence of our discovery and investigation of MAP catalysts.^{1–6,8} We expect that the results will continue to depend sensitively upon a combination of steric factors within each catalyst and substrate, and upon experimental conditions.

It should be noted that intermediate metallacyclobutanes in ruthenium-based metathesis catalysts that contain two chlorides⁹ are proposed to be 14 electron TBP species with the chlorides in axial positions.¹⁰ Axial chlorides would not be able to control the substitution pattern in the ruthenacyclobutane in the manner that we have achieved with the MAP complexes described here. No comparable *Z*-selectivities in homocoupling reactions have been reported for ruthenium-based catalysts.

Supplementary Material

Refer to Web version on PubMed Central for supplementary material.

Acknowledgments

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Table 1Homocoupling of 1-Octene (S_2).^a

Catalyst		time	%Conv.	%Z
2 _{Me}	Mo(NAd)(Me ₂ Pyr)(CHR ₂)(OHIPT) ^b	3 h	43	68
3 _{Mo}	Mo(NAr)(Pyr)(CHR ₂)(OHIPT) ^b	20 m	80	40
3 _W	W(NAr)(Pyr)(CHR ₂)(OHIPT) ^b	26 h	88	88
4 _W	W(NAr)(Pyr)(C ₃ H ₆)(OHIPT)	3 h	33	95
5 _{Mo}	Mo(NAr)(Pyr)(CHR ₂)(Mes ₂ Bitet) ^{*b}	15 m	58	70
5 _W	W(NAr)(Pyr)(CHR ₂)(Mes ₂ Bitet) ^{*b}	30 m	38	93
		2 h	72	88

^a 4 mol %; cat in C₆D₆ at 22 °C; R₂ = CMe₂Ph; Ad = 1-adamantyl; Ar = 2,6-i-Pr₂C₆H₃; OHIPT is the aryloxide shown in structure **2**; Mes₂Bitet is the ligand in structure **1** with R' = mesityl.

^b Prepared *in situ*; see Supporting Information.

Table 2

Screening of Substrates S₃–S₉.^a

S _n	Catalyst	t(h)	%Conv.	%Z
S ₃	3_W	3	40	91
S ₃	4_W	14	52	94
S ₃	6_W	3	62	93
S ₄	7_{Mo}	3	33	98
S ₅	8_W	1.5	69	92
S ₆	9_W	3	33	90
S ₇	3_W	3	30	94
S ₇	10_W	1.5	70	96
S ₇	7_{Mo}	3	33	98
S ₈	11_W	1.5	24	98
S ₉	11_W	3	52	98

^a 4 mol % cat in C₆D₆ at 22 °C in NMR tube; R₂ = CMe₂Ph; Ar^{Cl} = 2,6-Cl₂C₆H₃; Ar⁺ = 2,6-Me₂C₆H₃; Br₂Bitet is the ligand in structure **1** with R" = Br; Mes₂BitetOMe is the methyl-protected analog of the ligand in structure **1** with R" = Mesityl; R₃ = *t*-Bu.

^b Prepared *in situ*.

Table 3

Effect of Reduced Pressure on Z-Content (1 mol % cat).^a

Substrate	Catalyst	P(mmHg)	time(h)	%Conv.	%Z
S ₅	8_W	~0.5	0.2 (15)	25(>98)	>98(>98) ^b
S ₅	8_W (2%)	~760	1.5 (15)	84 (86)	97 (88)
S ₅	12_{Mo} ^c	~0.5	0.6 (16)	36 (34)	61 (61)
S ₅	12_{Mo} ^c	~760	0.6 (16)	24 (24)	61 (59)
S ₅	4_W	~0.5	5 (21)	7 (22)	>98(>98)
S ₅	4_W	~760	5 (21)	10 (27)	>98(>98)
S ₅	3_{Mo}	10	19	62	88
S ₇	3_{Mo}	~760	19	42	90
		~0.5	2	64	94
		~760	2	52	96
S ₁₀	12_{Mo} (2%)	~0.5	14	70	95

^a Reaction scale ~200 mg, neat substrate, catalyst added as a solid.^b 86% yield after 15 h.^c **12Mo** = Mo(NAr)(Pyr)(CHR₂)(Mes₂BitetOMe).

Table 4

Reactions Carried out at Elevated Temperatures.^a

Substrate	Catalyst	%	time(h)	°C(Bath)	%Conv.	%Z	%Yield
S ₁	4 _W	0.4	48	80	72	95	58
S ₂	4 _W	0.4	24	120	94	86	78
S ₃	10 _W	0.2	3	120	>98	77	77
	4 _W	0.2	4	120	63	93	56
S ₄	10 _W	0.2	24	100	94	88	65
	13 _W ^b	0.2	18	90	28	86	26
S ₅	13 _W ^b	2	23	100	>97	95	80
S ₆	13 _W ^b	1	16	100	97	87	80
S ₇	4 _W	0.2	1	100	46	91	42
	10 _W	0.2	18	100	74	94	90
S ₈	5 _W	4	24	100	46	>98	42
S ₉	4 _W	4	18	100	95	91	90
	10 _W	4	24	90	50	94	36

^a Reaction scale ~0.5 g to ~4 g; catalyst was dissolved in ~1 mL of benzene and substrate was added in one portion. The mixture was refluxed under N₂.^b W(NAr)(Py)(C₃H₆(OHIPT)).