

## MIT Open Access Articles

*Photodegradation mechanisms of tetraphenyl butadiene coatings for liquid argon detectors*

The MIT Faculty has made this article openly available. **Please share** how this access benefits you. Your story matters.

**Citation:** Jones, B J P, J K VanGemert, J M Conrad, and A Pla-Dalmau. "Photodegradation Mechanisms of Tetraphenyl Butadiene Coatings for Liquid Argon Detectors." *Journal of Instrumentation* 8, no. 01 (January 23, 2013): P01013–P01013.

**As Published:** <http://dx.doi.org/10.1088/1748-0221/8/01/p01013>

**Publisher:** Institute of Physics Publishing

**Persistent URL:** <http://hdl.handle.net/1721.1/88282>

**Version:** Original manuscript: author's manuscript prior to formal peer review

**Terms of use:** Creative Commons Attribution-Noncommercial-Share Alike



# Photodegradation Mechanisms of Tetraphenyl Butadiene Coatings for Liquid Argon Detectors

---

**B.J.P. Jones<sup>a\*</sup>, J.K. VanGemert<sup>b</sup>, J.M. Conrad<sup>a</sup>, A. Pla-Dalmau<sup>b</sup>.**

<sup>a</sup>*Massachusetts Institute of Technology,*

*77 Massachusetts Avenue, Cambridge, MA 02139, United States of America*

<sup>b</sup>*Fermi National Accelerator Laboratory,*

*PO Box 500, Batavia, IL 60510, United States of America*

*E-mail: bjpjones@mit.edu*

**ABSTRACT:** We report on studies of degradation mechanisms of tetraphenyl butadiene (TPB) coatings of the type used in neutrino and dark matter liquid argon experiments. Using gas chromatography coupled to mass spectrometry we have detected the ultraviolet-blocking impurity benzophenone (BP). We monitored the drop in performance and increase of benzophenone concentration in TPB plates with exposure to ultraviolet (UV) light, and demonstrate the correlation between these two variables. Based on the presence and initially exponential increase in the concentration of benzophenone observed, we propose that TPB degradation is a free radical-mediated photooxidation reaction, which is subsequently confirmed by displaying delayed degradation using a free radical inhibitor. Finally we show that the performance of wavelength-shifting coatings of the type envisioned for the LBNE experiment can be improved by 10-20%, with significantly delayed UV degradation, by using a 20% admixture of 4-tert-Butylcatechol.

**KEYWORDS:** Noble-liquid detectors; Photon detectors for UV, visible and IR photons; Scintillators, scintillation and light emission processes.

---

\*Corresponding Author

---

## Contents

<b>1. The Use of TPB In Neutrino and Dark Matter Experiments</b>	<b>1</b>
<b>2. Degradation Behavior of TPB</b>	<b>3</b>
<b>3. Identification of Impurities in TPB Powders With GCMS</b>	<b>3</b>
3.1 Use of GCMS to Identify and Quantify Impurities	3
3.2 Impurity GC Peaks in TPB Powders	4
<b>4. Monitoring Benzophenone Buildup in TPB Coatings</b>	<b>5</b>
<b>5. Benzophenone Spiking Studies, Active Coating Thickness and Discoloration</b>	<b>6</b>
<b>6. Chain Termination with Free Radical Inhibitors</b>	<b>8</b>
<b>7. Conclusions</b>	<b>9</b>

---

## 1. The Use of TPB In Neutrino and Dark Matter Experiments

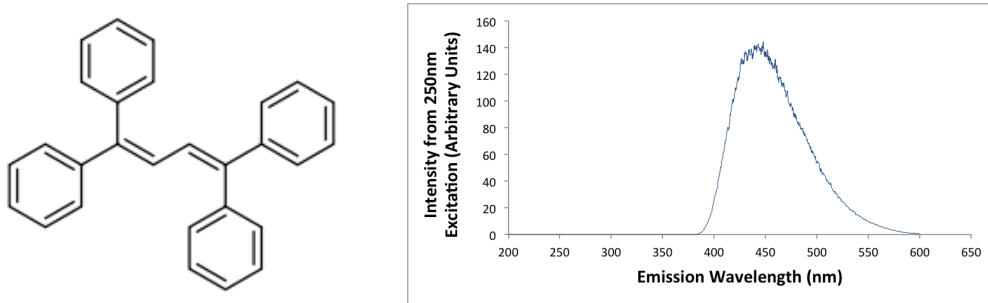
Liquid argon is the active medium in many existing and forthcoming particle physics detectors. Examples include liquid argon time projection chambers (LArTPCs) for neutrino studies, such as the ICARUS [1], ArgoNeuT [2], MicroBooNE [3], and LBNE [4] experiments, and low-background scintillation detectors for dark matter searches such as the DEAP [5] and MiniCLEAN [6] experiments.

Being inexpensive and inert, argon makes an ideal active medium for a LArTPC, where a long free-electron lifetime must be achieved. Argon is also an excellent scintillator, producing several tens of thousands of photons per MeV deposited. Further, high-purity argon is transparent to its own scintillation light. These characteristics make argon one of the favorite active media for low-background scintillation detectors.

Many LArTPCs also take advantage of argon scintillation light, which brings several benefits above simply recording charge deposits. Light collection systems which detect the flash of light accompanying a neutrino event can be used to acquire timing information about the event with a much higher precision than is possible with charge measurements alone. This timing information is an important background rejection tool for experiments deployed in pulsed neutrino beams, and is of particular importance for surface-based detectors such as MicroBooNE and LBNE, which both plan to implement extensive light collection systems. A light collection system in a LArTPC can cover a large detector volume with a relatively small number of channels, making the formation of trigger logic more straightforward than it would be using TPC wires, which typically number in the tens of thousands. There are also opportunities to use scintillation light for event reconstruction,

especially given its widely reported nontrivial time structure, which is used to determine the local ionization density of keV scattering events in dark matter detectors [7].

Scintillation light in liquid argon is produced at a wavelength of 128 nm. Argon is highly transparent at this wavelength, so the light can be collected by detectors outside of the active region. However, most optical detectors, including cryogenic photomultiplier tubes (PMTs), are insensitive to 128 nm light. Tetraphenyl butadiene (TPB) is a fluorescent compound with the chemical structure shown in Figure 1, which has a high absorption coefficient at 128 nm and a visible emission spectrum with a maximum around 450 nm. Since this is a good match to the peak quantum efficiency of standard cryogenic PMTs, TPB is commonly used as a wavelength shifter in liquid argon scintillation detectors. Thus, in order to make PMTs sensitive to argon scintillation light, a coating of TPB is applied either to the PMT face, as in ICARUS, or to a separate plate which sits in front of the PMT, which is the strategy favored by MicroBooNE, as well as MiniCLEAN. A lightguide-based detector which transmits TPB-shifted light to PMTs several meters from the interaction point is also being developed, for the future LBNE detector [8].



**Figure 1.** The chemical structure of TPB (left) and emission spectrum (right) [8]

There are several ways of applying TPB coatings, including evaporative coating, dip coating and painting. Evaporative coatings of pure TPB have the highest reported efficiency at 128 nm, with 1.3 visible photons being produced for each one incident 128 nm photon [9]. These coatings are fragile, however. In order to produce a more sturdy coating, TPB is often mixed into a toluene / polystyrene solution and painted onto the surface. In this paper we consider two types of films, which are similar to the coating planned for the MicroBooNE optical system and one coating being considered for the LBNE optical system, respectively. The MicroBooNE-type coating is designed to give the highest light yield possible for a polystyrene-based coating, and uses a 1:1 TPB-to-polystyrene ratio, with 1 g of each dissolved in 50 ml of toluene which is brush coated onto acrylic plates. At this concentration, the TPB crystallizes out of solution onto the plate surface as the toluene evaporates, leading to an opaque, rough finish. This coating will be referred to as the “rough” coating. The LBNE-type coating is designed to give a smooth optical surface for application on lightguides. For this coating, a 1:2 TPB-to-polystyrene ratio is used, with 0.5 g of TPB per 50 ml of toluene. With this composition, the TPB stays in solution as the coating dries, leading to a transparent polystyrene film with TPB suspended in the polystyrene matrix. This coating will be referred to as the “smooth” coating.

A third, polystyrene-free coating was prepared for GCMS studies, as polystyrene solutions were found to be damaging to the chromatography column. This coating consists of 1g TPB per

50 ml toluene brush coated onto the plate surface. All coated plates used in this study were cut from acrylic sheets of thickness 1/8" and were of size 4 cm x 4 cm. Half a milliliter of one of the aforementioned solutions was dripped onto the plate surface and then spread and smoothed with an acid brush. All plates were dried for at least 30 minutes in a fume hood before use.

## **2. Degradation Behavior of TPB**

It has been previously reported that TPB coatings degrade when exposed to UV light, both losing efficiency [10] and turning yellow in color [11] as a function of exposure. This UV sensitivity is an undesirable feature in light collection systems for large experiments, since it leads to more difficult installation procedures and comes with the risk of in situ degradation during the lifetime of the detector. Gaining understanding of and seeking strategies to prevent such degradation is therefore of significant interest in the field of argon scintillation detection, and may be of value to a wider community of TPB users.

In this paper, we measure the wavelength-shifting performance of TPB coatings with a Hitachi F-4500 fluorescence spectrophotometer, using an incident wavelength of 270 nm and measuring the emitted intensity at the TPB emission peak. This strategy prevents backgrounds from direct reflections or ambient light from contributing significantly to the measured efficiency. Using a reference plate kept in the dark we can correct for spectrophotometer drift, which leads to changes of a few percent in the reported efficiency over the duration of any study.

Two grades of TPB are commercially available: scintillation grade, which has a purity specification of >99.9%, and standard grade, which has a purity specification of >99%. It has been noted before that the standard grade product is much more susceptible to discoloration, turning bright yellow after a few hours of exposure to sunlight. The scintillation grade product does not turn bright yellow, rather an off-white color, and has a longer characteristic degradation time.

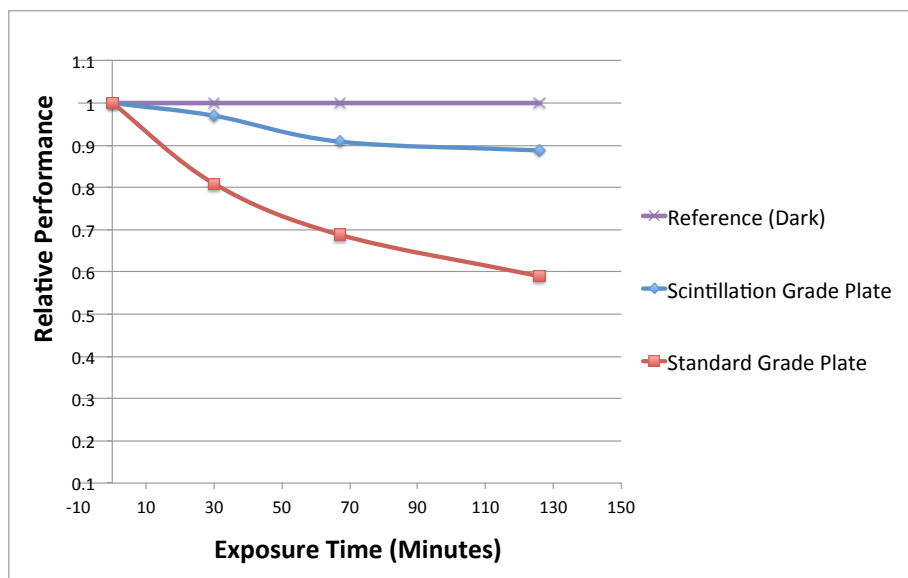
The relative efficiency of plates produced using the "rough" coating with standard and scintillation grade TPB as a function of exposure to sunlight is shown in Figure 2. Both curves are normalized to have their initial efficiency be equal to 1, and each data point is corrected for spectrophotometer drift using the measured efficiency of the reference plate.

The intensity of the sunlight naturally varied over the course of the study, so the scale on the horizontal axis should not be taken to be linear. However by using sunlight we can ensure that both plates received the same exposure, which has proved difficult to implement with lamp-based studies, where slight changes to the geometry of the setup or stability of the lamp are difficult to control. In this study, the standard grade plate is seen to degrade significantly faster than the scintillation grade plate. The observation that the standard grade sample both loses performance and experiences discoloration at a higher rate than the scintillation grade sample is suggestive that both of these features are a consequence of initial impurities present in the sample.

## **3. Identification of Impurities in TPB Powders With GCMS**

### **3.1 Use of GCMS to Identify and Quantify Impurities**

We used an Agilent 7820A Gas Chromatograph with a 5975 Mass Selective Detector (GCMS) to identify and quantify impurities in TPB powders and coatings. This process involves two stages.



**Figure 2.** Degradation of standard and scintillation grade TPB plates with “rough” coating, as a function of exposure to direct sunlight

First, different components of the sample are separated by their diffusion time in a gas chromatography (GC) column. Then, each chromatographically separated component is fed into a mass spectrometer (MS). The mass spectrum can be compared to a database of known spectra to elucidate a likely identity for the component.

After making preliminary mass-scan runs which identified impurities present in the sample, a custom GC method was created to accurately quantify their concentrations. This method incorporated a temperature gradient of 120°C to 250°C, increasing at a rate of 5°C per minute. This profile was chosen due to the different retention times exhibited by benzophenone and TPB. Benzophenone would elute with the solvent at higher initial temperatures, whereas TPB would remain in the chromatography column at lower oven temperatures. The method was also set to maintain a constant flow rather than a constant pressure. All injections were manual with a 5 $\mu$ l syringe which was rinsed with solvent between each injection.

The MS was set to acquire data in selective ion mode (SIM) because of the sensitivity needed to detect the low benzophenone levels. Each ion group was set to monitor the four most common ions for each compound. Due to the small size of the benzophenone peak compared to the TPB signal, in most cases a manual integration was required.

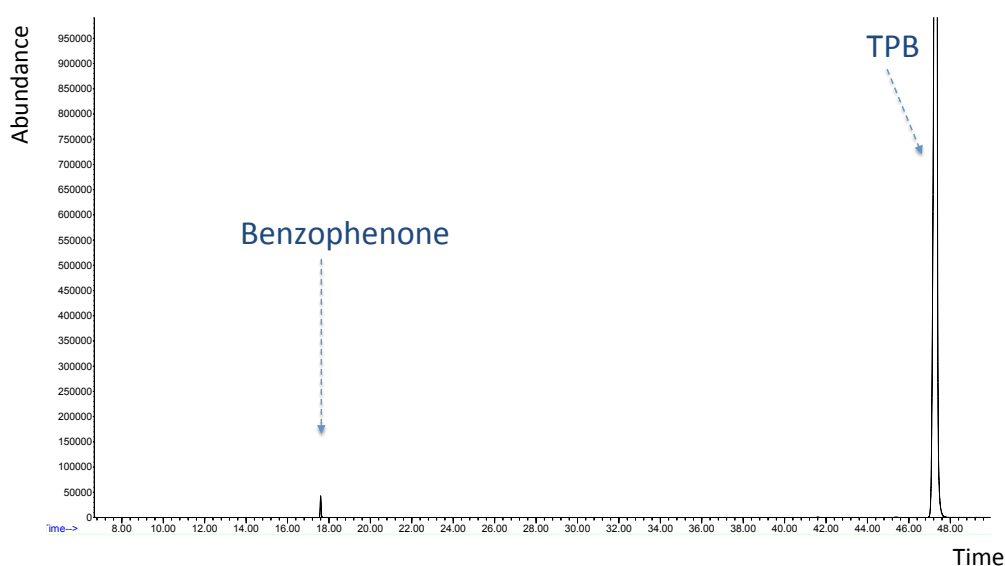
### 3.2 Impurity GC Peaks in TPB Powders

An example chromatograph for standard grade TPB dissolved in toluene is shown in Figure 3. The large peak at 47 minutes is identified as TPB. The peak at around 17.5 minutes is identified with a high confidence level as benzophenone, which is a commercially available UV blocker and photoinitiator [12] and a likely photooxidation product of TPB. The benzophenone concentration was measured by comparing the BP to TPB peak areas in three powder samples: unexposed standard grade, unexposed scintillation grade, and standard grade TPB exposed to UV light to the point of

Sample	Benzophenone / TPB Peak Ratio
Scintillation Grade (Unexposed, White)	$4.13 \times 10^{-6}$
Standard Grade (Unexposed, White)	$8.43 \times 10^{-6}$
Standard Grade (Exposed, Yellow)	$1.82 \times 10^{-5}$

**Table 1.** Benzophenone-to-TPB peak area ratios in powder TPB samples.

yellowing. The benzophenone-to-TPB peak ratio from each are reported in Table 1



**Figure 3.** Chromatograph of unexposed standard grade TPB in toluene.

When TPB is dissolved in toluene and applied as a plate coating, the benzophenone concentration is seen to increase significantly, with the BP-to-TPB peak ratio increasing from order  $10^{-4}\%$  in power samples, to around 0.01% in samples taken from freshly coated plates. This may be a result of the necessary light exposure involved in painting and drying the plate, or of a reaction which occurs in solution. The initial BP concentration varies between batches of plates, but sets made from a common solution and subjected to a common storage procedure appear to have consistent benzophenone concentrations to within 20%.

#### 4. Monitoring Benzophenone Buildup in TPB Coatings

The results of Section 3 are suggestive that the buildup of the benzophenone occurs when TPB is exposed to UV light. To test this hypothesis we prepared many identical plates with scintillation grade TPB coatings. TPB was dissolved in toluene to make a solution of concentration 0.02 g/ml, and each plate received one coating made from 0.5 ml of this solution. As mentioned in Section 1, polystyrene was omitted for this test because it was found that polystyrene solutions would not give consistent results and eventually lead to clogging of the GC column.

Each plate was stored in the dark, wrapped in a mylar sheet, until we were ready to begin exposing it to UV light. Immediately after removing each plate from the dark we measured the initial wavelength shifting efficiency using the spectrophotometer procedure described in Section 2. It was then exposed to a 365nm UV lamp for a number of hours, and then the efficiency was measured again. The ratio of these two measurements tells us the extent of the plate degradation. Next, the plate was placed into a small bath of 20 ml toluene and sonicated for five minutes in order to dissolve off the TPB coating. Two  $\mu\text{l}$  the solution was injected into the spectrometer and the benzophenone / TPB peak ratio was recorded. Using many plates we monitor the benzophenone content and plate performance as a function of exposure time, which is shown in Figure 4. Since several plates were under the lamp simultaneously, the geometry of the setup means that not all got exactly equal exposure. This leads to some uncertainty in the total integrated UV exposure, which is evident in the horizontal scatter observed in the data points of Figure 4. Plotting benzophenone content against plate performance loss bypasses this issue and yields a very clear correlation, as is shown in Figure 5. We assign a constant error of 7% in absolute performance to the spectrophotometer measurements, which is determined by the spread of efficiencies measured for the unexposed reference plate, and 10% to the benzophenone / TPB peak ratio, which is the approximate spread observed in repeated measurements of the same solution.

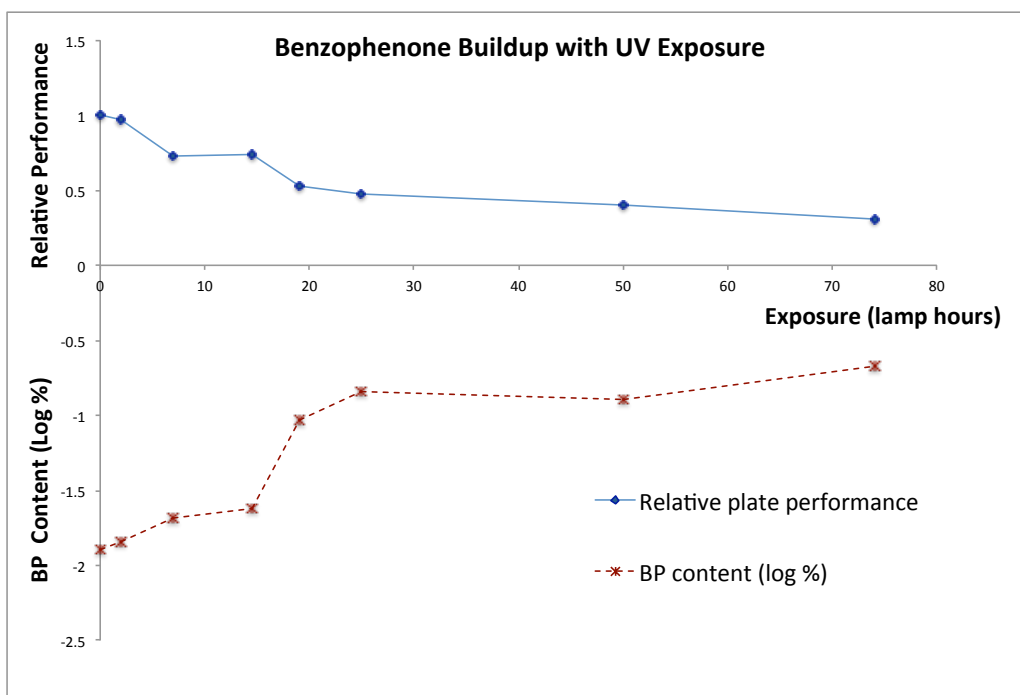
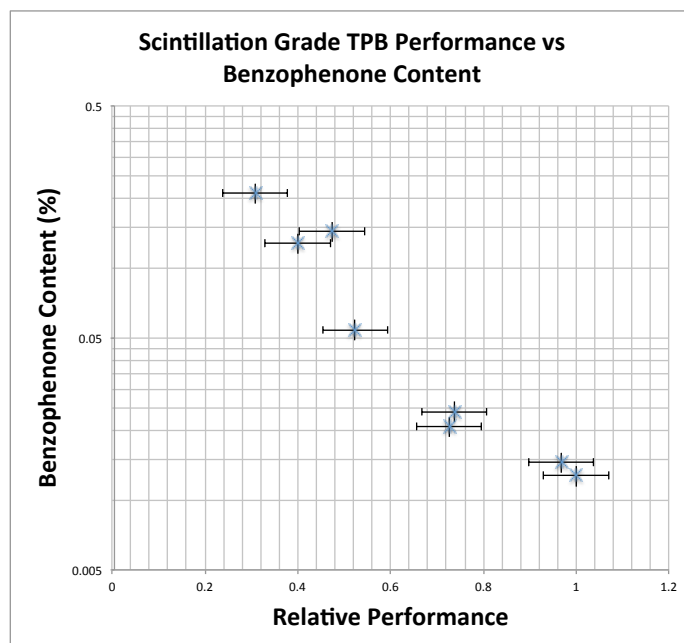


Figure 4. Degradation of plate performance and accumulation of benzophenone with exposure time

## 5. Benzophenone Spiking Studies, Active Coating Thickness and Discoloration

We have observed the accumulation of sub-percent levels of benzophenone in TPB coatings under exposure to UV light. Benzophenone is a commercially available UV blocker and photoinitiator,





**Figure 5.** Performance vs benzophenone content (note log scale)

so it is tempting to attribute the loss of performance of TPB coatings to the partial absorption of incident light by the accumulated benzophenone, preventing it from being absorbed and re-emitted by TPB.

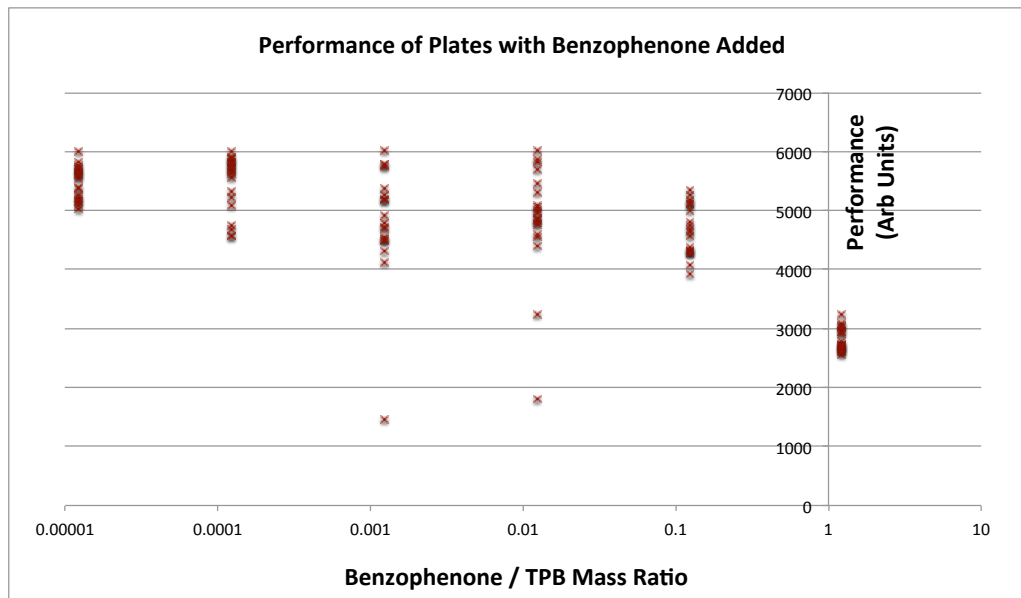
In order to test the UV-blocking properties of benzophenone built up within the TPB layer, we prepared “smooth” plates spiked with different benzophenone concentrations and looked for a change in performance relative to unspiked plates. In order to ensure that we were not misled by natural plate-to-plate variations, we acquired high statistics by preparing four plates at each concentration and measured the performance of each plate at four different locations on its surface. The results of this study are shown in Figure 6.

We note first that up to a benzophenone concentration of around 10% by weight, no significant loss of performance is observed, except for a few outlying points which seem to indicate a tendency for a less uniform TPB distribution within the coating when benzophenone is added. Only when there is twice as much BP as TPB do we see the levels of degradation observed in UV exposure studies.

In TPB-coated plates, only the very top layer of the coating under study is optically active, with all UV light being absorbed in a thin surface layer. In this case, the local benzophenone density at the plate surface would be expected to be much higher than the global benzophenone concentration across the entire coating, since it has received a higher UV light exposure. This produces an effective UV-blocking film of benzophenone which prevents the underlying TPB from receiving the incident light.

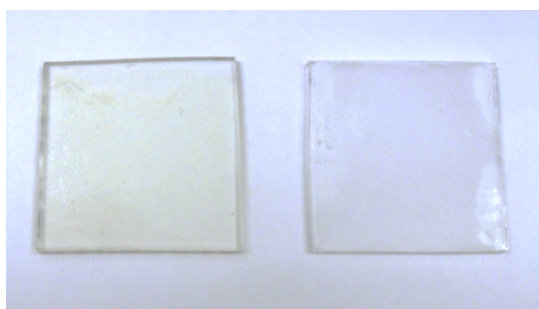
This type of skin effect is common in polymer photodegradation [13], and is to be expected in TPB, especially in light of recent studies which demonstrate that the efficiency of evaporative coatings is independent of their thickness [14].

Our spiking results suggest that the skin layer of degraded plates must have a composition of almost 100% benzophenone, and from this fact and the known density of polystyrene we can estimate the thickness of the optically active TPB layer to be of the order 100 nm, from a total coating thickness of order 100  $\mu\text{m}$ .



**Figure 6.** Performance of Benzophenone Spiked Plates

We also note an increased tendency for plates made with added benzophenone to turn yellow in color under light exposure. A photograph showing two similarly degraded plates, prepared with and without added benzophenone, is shown in Figure 7. We speculate that this effect may be due to the buildup of benzophenone derivatives, many of which are known to be both yellow in color and strongly UV-blocking [15]



**Figure 7.** Similarly degraded plates with (left) and without (right) added benzophenone

## 6. Chain Termination with Free Radical Inhibitors

The amount of benzophenone accumulation displayed in Figure 4 is initially approximately exponential as a function of light exposure. This is suggestive of a free radical-mediated chain reac-

tion, which is a common mechanism of light-induced degradation, often encountered in polymer chemistry. [16]. Several compounds exist which can be used to terminate such chain reactions prematurely, and thus prevent the light-generated free radicals from causing further degradation. To test whether this degradation is indeed a free radical-mediated photooxidation, and to see if it is possible to stabilize the coating, we investigated two such chain termination compounds: 3,5-di-tert-butyl-4-hydroxytoluene (BHT) and 4-tert-Butylcatechol (BC).

For both "rough" and "smooth" coatings, BHT was found to produce a less uniform coating with regions of high and low TPB density. For this reason, it was difficult to extract consistent results from coatings with BHT admixtures. We are continuing to investigate methods of improving the coating quality of BHT / TPB mixtures.

BC admixtures did produce a homogeneous coating. We even observed a slight tendency towards more TPB remaining suspended in the polystyrene matrix rather than crystallizing on the surface for the "rough" plates.

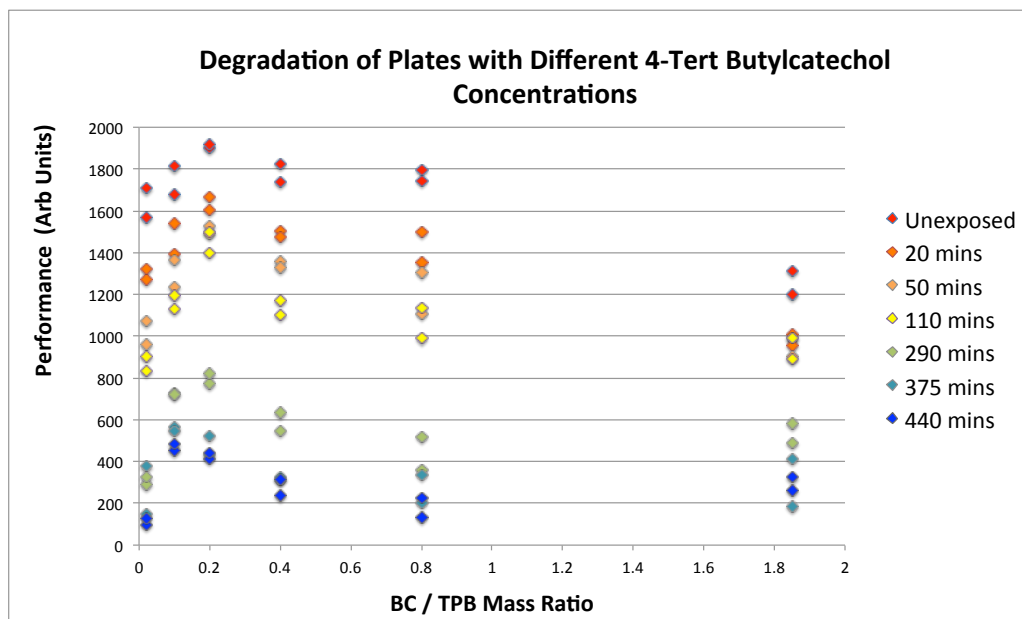
We investigated plates with BC / TPB ratios between 0.1 ppm and 200% by weight, and a stabilizing effect was observed for mass ratios above 1% BC / TPB. After establishing the relevant concentration range, we prepared many plates with mass ratios between 1 and 200% and monitored their degradation over a few-hour period of simultaneous exposure to direct sunlight using the previously discussed spectrophotometer procedure. Figure 8 shows the absolute efficiencies of plates prepared at various concentrations, as a function of exposure time. At each concentration point we used two identical, separately prepared plates, to illustrate the spread in coating performance. The plates used were made with the "smooth" coating. The results show that the addition of around 20% of BC appears not only to improve the initial performance of the coating, but also to slow its degradation. This is a confirmation that we are indeed seeing a free radical-mediated reaction.

We tested that this effect is repeatable by preparing many plates with 0% and 20% BC, in both the "rough" and "smooth" styles, and monitoring their performance as a function of sunlight exposure. The data for this study are shown in Figure 9. We note that the "smooth" plates reliably see a 10-20% improvement in initial light yield and an extended degradation time, such that after 200 minutes of exposure, the stabilized coatings are ~200% as efficient at the nonstabilized coatings.

The "rough" plates are not improved. This appears to be due to the fact that less TPB crystallizes onto the surface when BC is added, with more remaining suspended in the polystyrene layer. It may be possible to turn this effect into an overall improvement for the "rough" plates also, by increasing the quantity of TPB in the mixture to restore the surface density to its value when no BC is added. We plan to investigate this possibility in the near future. Since the surface layer is responsible for most of the wavelength-shifting capability of the "rough" plates, and most of the BC remains in the polystyrene matrix, we see little effect upon the coating lifetime for this sample.

## 7. Conclusions

We have identified benzophenone as a compound produced in TPB films which are exposed to ultraviolet light. Since benzophenone is a known ultraviolet blocker and photoinitiator [12], it seems likely that its presence is intimately related to the performance degradation observed in TPB coatings. Using GCMS and fluorescence measurements we have found a strong correlation between benzophenone concentration and performance loss in TPB coatings with various UV exposures.



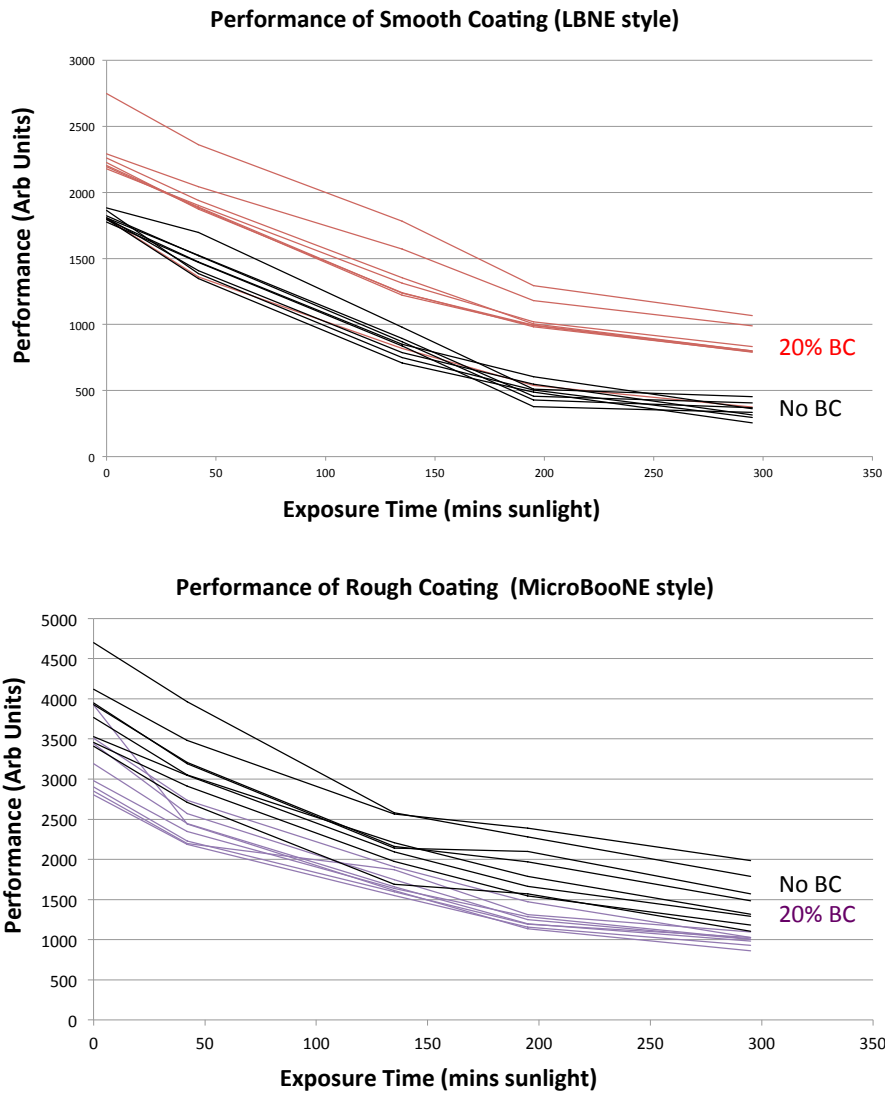
**Figure 8.** Degradation of plates with different BC / TPB mass ratios

By preparing plates with BP / TPB mixtures we have shown that if benzophenone is the primary cause of performance loss, it must be accumulated within a thin, optically active surface layer of the coating. This type of skin effect is common in photochemistry [13] and is supported by the results of other groups who have studied TPB films of varying thicknesses [14].

The previously reported coating discoloration [11] also appears to be related to the presence of benzophenone, as plates spiked with additional benzophenone are observed to turn significantly more yellow than TPB-only plates under exposure to similar light levels.

The initially exponential buildup of benzophenone under light exposure is suggestive of a free radical-mediated TPB photooxidation reaction, which was tested by exposing plates with an admixture of the free radical inhibitor 4-tert-Butylcatechol. The delay in degradation provided by BC confirms that TPB degradation is a free radical-mediated reaction.

An admixture of 20% BC by weight leads to an initial improvement of the coating efficiency by 10-20% and a delayed fluorescence degradation such that after 200 minutes of sunlight exposure, the BC-protected coating is more than 200% as efficient as the unprotected coating. A program to study other stabilizing compounds is underway. It is possible that other free radical inhibitors may be found which can improve the lifetime of TPB coatings even more significantly.



**Figure 9.** Study of 20% vs 0% BC plates, with both the "rough" and "smooth" style coating

## Acknowledgments

We would like to thank Christie Chiu for her help in performing preliminary studies related to this paper. The authors thank the National Science Foundation (NSF-PHY-084784) and Department Of Energy (DE-FG02-91ER40661). This work was supported by the Fermi National Accelerator Laboratory, which is operated by the Fermi Research Alliance, LLC under Contract No. De-AC02-07CH11359 with the United States Department of Energy.

## References

- [1] **ICARUS Collaboration** Collaboration, A. Menegolli, *ICARUS and status of liquid argon technology*, *J.Phys.Conf.Ser.* **375** (2012) 042057.
- [2] **ArgoNeuT Collaboration** Collaboration, O. Palamara, *Neutrino Detection in the ArgoNeuT LAr TPC*, arXiv:1110.3070.
- [3] B. J. P. Jones, *The Status of the MicroBooNE Experiment*, *PoS EPS-HEP2011* (2011) 436, [arXiv:1110.1678].
- [4] **LBNE Collaboration** Collaboration, T. Akiri et al., *The 2010 Interim Report of the Long-Baseline Neutrino Experiment Collaboration Physics Working Groups*, arXiv:1110.6249.
- [5] **DEAP Collaboration** Collaboration, M. Boulay, *DEAP-3600 Dark Matter Search at SNOLAB*, *J.Phys.Conf.Ser.* **375** (2012) 012027, [arXiv:1203.0604].
- [6] **MiniCLEAN Collaboration** Collaboration, K. Rielage, *Status and prospects of the MiniCLEAN dark matter experiment*, *AIP Conf.Proc.* **1441** (2012) 518–520.
- [7] P. Kryczynski, *Pulse Shape Discrimination in liquid argon and its implications for Dark Matter searches using depleted argon*, *Acta Phys.Polon.* **B43** (2012) 1509, [arXiv:1210.1019].
- [8] B. Baptista, L. Bugel, C. Chiu, J. Conrad, C. Ignarra, et al., *Benchmarking TPB-coated Light Guides for Liquid Argon TPC Light Detection Systems*, arXiv:1210.3793.
- [9] V. Gehman, S. Seibert, K. Rielage, A. Hime, Y. Sun, et al., *Fluorescence Efficiency and Visible Re-emission Spectrum of Tetraphenyl Butadiene Films at Extreme Ultraviolet Wavelengths*, *Nucl.Instrum.Meth.* **A654** (2011) 116–121, [arXiv:1104.3259].
- [10] C. Chiu, C. Ignarra, L. Bugel, H. Chen, J. Conrad, et al., *Environmental Effects on TPB Wavelength-Shifting Coatings*, *JINST* **7** (2012) P07007, [arXiv:1204.5762].
- [11] R. Jerry, L. Winslow, L. Bugel, and J. Conrad, *A Study of the Fluorescence Response of Tetraphenyl-Butadiene*, arXiv:1001.4214.
- [12] M. Baril et al., “Iarc monographs on the evaluation of carcinogenic risks to humans - benzophenone.”
- [13] N. Grassie and G. Scott, *Polymer Degradation Stabilisation*. Press Syndicate of University of Cambridge, 1985.
- [14] R. Francini et al., *Studies of thin film tpb coatings*, October, 2012. Private communication.
- [15] Sigma-Aldrich, *Materials safety data sheets: products 92841, 59647, 413151, h36206, .*
- [16] H. R. Alcock, F. W. Lampe, and J. E. Mark, *Contemporary Polymer Chemistry*. Prentice Hall, 2003.